

Gases Study Guide**Test Format**

Multiple Choice
Gas Law Problems
Gas Stoich Problems
Short Answer

Equations & Conversions

$$P = \frac{F}{A}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$PV = nRT$$

$$R = 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

$$1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$$

$$\frac{\text{Rate of Effusion of L}}{\text{Rate of Effusion of H}} = \sqrt{\frac{\text{Molar Mass of H}}{\text{Molar Mass of L}}}$$

$$K = ^\circ\text{C} + 273 \quad 0.001 \text{ kL} = 1 \text{ L} = 1,000 \text{ mL}$$

$$\text{At STP, 1 mole} = 22.4 \text{ L} \quad \text{KHDSdcm}$$

1. A box weighs 180 lbs and has a base that is 6 inches by 9 inches. What is the pressure this box exerts on the surface it is sitting on?

$$P = \frac{F}{A} \quad \frac{180 \text{ lbs}}{54 \text{ in}^2} = \boxed{3.3 \text{ psi}}$$

$$54 \text{ in}^2 = \text{area}$$

2. What is atmospheric pressure?

The pressure exerted by the collisions of gases in the air surrounding Earth (N_2 , O_2 , etc).

3. Convert 1.088 atm to:

a. mm Hg $\frac{1.088 \text{ atm}}{1} \cdot \frac{760 \text{ mmHg}}{1 \text{ atm}} = \boxed{826.9 \text{ mmHg}}$

b. kPa $\frac{1.088 \text{ atm}}{1} \cdot \frac{101.3 \text{ kPa}}{1 \text{ atm}} = \boxed{110.2 \text{ kPa}}$

4. What is a barometer and who invented it?

A tool that measures room/air pressure. Invented by Torricelli.



5. How are temperature and kinetic energy related?

Direct. As temp \uparrow , KE \uparrow , molecules move faster.
As temp \downarrow , KE \downarrow , molecules move slower.

6. Briefly describe the five assumptions of the kinetic-molecular theory.

Gases have a lot of tiny particles that are always in motion that is random, fast, and continuous. Particles are far apart without any forces of attraction, so their collisions are elastic \rightarrow bouncing off of each other/walls. The speed of the molecules change with temperature.

Chap13 Gases Study Guide Chemistry Answers

Paul B. Kelter



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Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall, Donald J. DeCoste, 2000

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1. Understanding the eBook Chap13 Gases Study Guide Chemistry Answers
 - The Rise of Digital Reading Chap13 Gases Study Guide Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chap13 Gases Study Guide Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chap13 Gases Study Guide Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chap13 Gases Study Guide Chemistry Answers
 - Personalized Recommendations
 - Chap13 Gases Study Guide Chemistry Answers User Reviews and Ratings
 - Chap13 Gases Study Guide Chemistry Answers and Bestseller Lists
5. Accessing Chap13 Gases Study Guide Chemistry Answers Free and Paid eBooks
 - Chap13 Gases Study Guide Chemistry Answers Public Domain eBooks
 - Chap13 Gases Study Guide Chemistry Answers eBook Subscription Services
 - Chap13 Gases Study Guide Chemistry Answers Budget-Friendly Options

6. Navigating Chap13 Gases Study Guide Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chap13 Gases Study Guide Chemistry Answers Compatibility with Devices
 - Chap13 Gases Study Guide Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chap13 Gases Study Guide Chemistry Answers
 - Highlighting and Note-Taking Chap13 Gases Study Guide Chemistry Answers
 - Interactive Elements Chap13 Gases Study Guide Chemistry Answers
8. Staying Engaged with Chap13 Gases Study Guide Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chap13 Gases Study Guide Chemistry Answers
9. Balancing eBooks and Physical Books Chap13 Gases Study Guide Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chap13 Gases Study Guide Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chap13 Gases Study Guide Chemistry Answers
 - Setting Reading Goals Chap13 Gases Study Guide Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chap13 Gases Study Guide Chemistry Answers
 - Fact-Checking eBook Content of Chap13 Gases Study Guide Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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