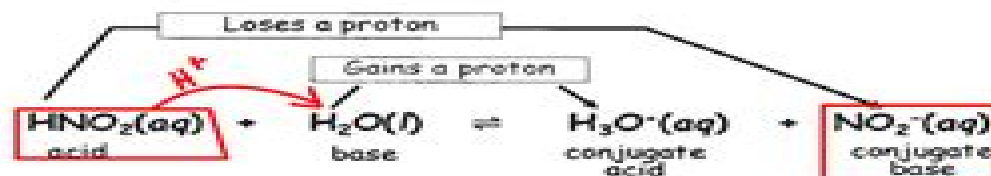


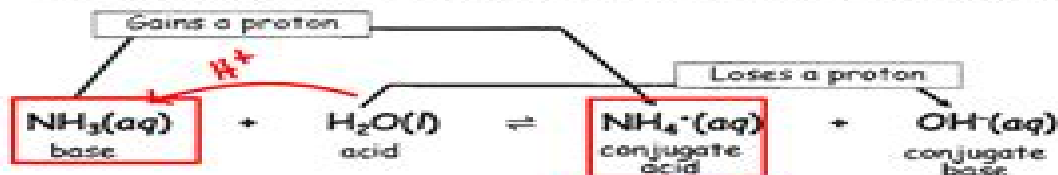
**B. Conjugate Acid-Base Pairs**

1. Every acid has a conjugate base and every base has a conjugate acid.
2. A conjugate acid-base pair consists of two species that differ from each other by just one proton,  $H^+$ .
3. When a Brønsted acid donates a proton, what remains of the acid is known as a *conjugate base*.



4. The two species  $\text{HNO}_2$  and  $\text{NO}_2^-$  are known as a *conjugate acid-base pair* or simply a *conjugate pair*.

5. When a Brønsted base accepts a proton, the newly formed protonated species is known as a *conjugate acid*.



6. The two species  $\text{NH}_3$  and  $\text{NH}_4^+$  are known as a *conjugate acid-base pair* or simply a *conjugate pair*.

**PROBLEM-SOLVING CLASS ACTIVITY 1**

For the following reaction in aqueous solution, identify the conjugate acid-base pairs.



# Chapter 14 Chemistry Study Guide Answers

**Susan McMurry**



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