

Physics

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(Chapter 13) (Kinetic Theory) (Additional Exercises)
(Class 11)

Question 13.14:

Given below are densities of some solids and liquids. Give rough estimates of the size of their atoms:

Substance	Atomic Mass (u)	Density (10^3 Kg m^{-3})
Carbon (diamond)	12.01	2.22
Gold	197.00	19.32
Nitrogen (liquid)	14.01	1.00
Lithium	6.94	0.53
Fluorine (liquid)	19.00	1.14

[Hint: Assume the atoms to be 'tightly packed' in a solid or liquid phase, and use the known value of Avogadro's number. You should, however, not take the actual numbers you obtain for various atomic sizes too literally. Because of the crudeness of the tight packing approximation, the results only indicate that atomic sizes are in the range of a few Å].

Answer 13.14:

Substance	Radius (Å)
Carbon (diamond)	1.29
Gold	1.59
Nitrogen (liquid)	1.77
Lithium	1.73
Fluorine (liquid)	1.88

Atomic mass of a substance = M and Density of the substance = ρ

Avogadro's number = $N = 6.023 \times 10^{23}$

Volume of N number of molecules = $\frac{4}{3} \pi r^3 N$... (i)

Volume of one mole of a substance = $\frac{M}{\rho}$... (ii)

$$\frac{4}{3} \pi r^3 N = \frac{M}{\rho} \Rightarrow r = \sqrt[3]{\frac{3M}{4\pi\rho N}}$$

For gold: $M = 197.00 \times 10^{-3} \text{ kg}$ $\rho = 19.32 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 197 \times 10^{-3}}{4\pi \times 19.32 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.77 \text{ Å}$$

Hence, the radius of a liquid nitrogen atom is 1.77 Å.

For lithium: $M = 6.94 \times 10^{-3} \text{ kg}$ $\rho = 0.53 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 6.94 \times 10^{-3}}{4\pi \times 0.53 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.73 \text{ Å}$$

Hence, the radius of a lithium atom is 1.73 Å.

For liquid fluorine: $M = 19.00 \times 10^{-3} \text{ kg}$ and $\rho = 1.14 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 19 \times 10^{-3}}{4\pi \times 1.14 \times 10^3 \times 6.023 \times 10^{23}} \right)^{\frac{1}{3}} = 1.88 \text{ Å}$$

Hence, the radius of a liquid fluorine atom is 1.88 Å.

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**Ajawad I. Haija, M. Z. Numan, W. Larry
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1. Understanding the eBook Chapter 13 Physics Manual Solutions
 - The Rise of Digital Reading Chapter 13 Physics Manual Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 13 Physics Manual Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 13 Physics Manual Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 13 Physics Manual Solutions
 - Personalized Recommendations
 - Chapter 13 Physics Manual Solutions User Reviews and Ratings
 - Chapter 13 Physics Manual Solutions and Bestseller Lists
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 - Chapter 13 Physics Manual Solutions Enhanced eBook Features
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- Adjustable Fonts and Text Sizes of Chapter 13 Physics Manual Solutions
 - Highlighting and Note-Taking Chapter 13 Physics Manual Solutions
 - Interactive Elements Chapter 13 Physics Manual Solutions
8. Staying Engaged with Chapter 13 Physics Manual Solutions
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 13 Physics Manual Solutions
9. Balancing eBooks and Physical Books Chapter 13 Physics Manual Solutions
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 13 Physics Manual Solutions
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 13 Physics Manual Solutions
- Setting Reading Goals Chapter 13 Physics Manual Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 13 Physics Manual Solutions
- Fact-Checking eBook Content of Chapter 13 Physics Manual Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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



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