

CHAPTER 19 STUDY GUIDE FOR CONTENT MASTERY

Acids and Bases**Section 19.1 Acids and Bases: An Introduction**

In your notebook, read about the properties of acids and bases.

For each description below, write and *b* if the statement is a property of an acid or *b* if it is not a property of acids. If the property does not apply to either an acid or a base, write *n*.*Both* _____ 1. Can form from a proton transfer*Acid* _____ 2. Has a sour taste*Both* _____ 3. Contains more hydrogen ions than hydronium ions*Base* _____ 4. Has a bitter taste*Both* _____ 5. Reacts with acids*Neither* _____ 6. Has a high pH*Neither* _____ 7. Contains equal numbers of hydrogen and hydroxide ions*Base* _____ 8. Turns blue*Acid* _____ 9. Turns red

In your notebook, read about the different models of acids and bases.

Acid/base	Ionization theory	Octet rule acid/base
Acidic base	hydrogen ions	chloride acid hydroxide

The *(19)* **Arrhenius** model of acids and bases states that an acid contains *hydrogen* ions and a base contains *hydroxide* ions.In water, a base contains the *(19)* **hydroxide** group and dissociates to produce *(19)* **hydroxide** ions in aqueous solutions.According to the *(19)* **Arrhenius** model, an acid dissociates to form *(19)* **hydrogen** ions, and a base accepts *(19)* **hydrogen** ions.According to the model, in an aqueous solution, each acid has a *(19)* **conjugate base**, and each base has a *(19)* **conjugate acid**.

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Chemistry: Matter and Change

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CHAPTER 19 STUDY GUIDE FOR CONTENT MASTERY**Section 19.2 Strengths of Acids and Bases**

In your notebook, read about the strengths of acids.

Circle the letter of the choice that best completes the statement or answers the question.

- Acid A and Acid B are of equal concentrations and are mixed with a conductivity apparatus. When the electrodes are placed in acid A, the bulb glows more brightly. When they are placed in acid B, the bulb glows more brightly. Which of the following is true?
 - Acid A is stronger than acid B.
 - Acid B is stronger than acid A.
 - Acid A and acid B are of equal strength.
 - Strengths of strength can be read from the meter.
- A chemical equation for the ionization of an acid into a single ion in the right (\rightarrow) is written for review and problem sets of the chapter. Which of the following is not?
 - The anion does not indicate either strength.
 - The cation and anion are of equal strength.
 - The cation acid is half ionized.
 - The cation acid is weak.
- Acids are acids and bases are bases. What is true about the conjugate base?
 - Its conjugate base is explosive.
 - Its conjugate base is strong.
 - Its conjugate base is weak.
 - Its conjugate base is made by reducing the strength of the conjugate acid.
 - It is acidic, it is weak and produces
 - a mixture of anions and cations.
 - a base.
 - anions, but no hydroxide ions.
 - Why are *(19)* **water** and **methanol**?
 - The concentration of water does not affect the ionization.
 - The equilibrium is not stable.
 - The solution contains a high concentration of neutral acid molecules.
 - The solution contains a high concentration of neutral acid molecules.
 - Which of the following dissociations results in an acid and a hydroxide base in solution?
 - a strong acid
 - a strong base
 - a weak acid
 - a weak base
 - In general, compounds formed from acids and bases are
 - strong acids
 - strong bases
 - weak acids
 - weak bases

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John A. Olmsted, Gregory M. Williams

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