

Name: \_\_\_\_\_

Name: \_\_\_\_\_

**MULTIPLE CHOICE.** Choose the one alternative that best completes the statement or answers the question.

13. The total concentration of ions in a 0.250 M solution of  $\text{HCl}$  is \_\_\_\_\_.

- A) 0.250 M
- B) 0.500 M
- C) 1.25 M
- D) 0.750 M
- E) insolvability.

13. \_\_\_\_\_

14. A weak electrolyte consists primarily of an \_\_\_\_\_ in solution.

- A) cation
- B) anion
- C) solute
- D) water
- E) an isotope.

14. \_\_\_\_\_

15. Which of the following are strong electrolytes?

- A)  $\text{HCl}$
- B)  $\text{HCO}_3\text{H}_2\text{O}_2$
- C)  $\text{NH}_3$
- D)  $\text{KCl}$ 
  - A)  $\text{HCl}$ ,  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{NH}_3$ ,  $\text{KCl}$
  - B)  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{KCl}$
  - C)  $\text{HCl}$ ,  $\text{NH}_3$ ,  $\text{KCl}$
  - D)  $\text{HCl}$ ,  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{KCl}$
  - E)  $\text{HCl}$ ,  $\text{KCl}$

15. \_\_\_\_\_

16. Which of the following are weak electrolytes?

- A)  $\text{HCl}$
- B)  $\text{HCO}_3\text{H}_2\text{O}_2$
- C)  $\text{NH}_3$
- D)  $\text{KCl}$ 
  - A)  $\text{HCl}$ ,  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{NH}_3$ ,  $\text{KCl}$
  - B)  $\text{HCl}$ ,  $\text{KCl}$
  - C)  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{KCl}$
  - D)  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{NH}_3$
  - E)  $\text{HCl}$ ,  $\text{HCO}_3\text{H}_2\text{O}_2$ ,  $\text{KCl}$

16. \_\_\_\_\_

17. Which aqueous solution of \_\_\_\_\_ will precipitate  $\text{CaCO}_3$ .

- A)  $\text{NaCl}$  and  $\text{KBr}$
- B)  $\text{Li}_2\text{CO}_3$  and  $\text{CaCl}_2$
- C)  $\text{K}_2\text{SO}_4$  and  $\text{CaCl}_2$
- D)  $\text{NaOH}$  and  $\text{BaCl}_2$
- E)  $\text{NaBr}$  and  $\text{AgNO}_3$

17. \_\_\_\_\_

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**Kenneth W. Whitten, Raymond E. Davis**

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