

Chemistry Study Guide Chapter 4 - Key

1. Define the following terms: (a) atom, (b) atomic number, (c) mass number, (d) isotopes (give 2 definitions), (e) atomic mass unit (amu), (f) nucleus, (g) spectroscopy, (h) ground state, (i) excited state, (j) atomic mass, (k) quantum leap, (l) photon.

(a) Atom: the smallest particle of an element that still has the element's properties.

(b) Atomic number: the number of protons in the nucleus of an atom.

(c) Mass number: the number of protons plus the number of neutrons in an atom.

(d) Isotopes: Atoms with the same atomic number but different mass numbers or atoms with the same number of protons but a different number of neutrons.

(e) Atomic mass unit (amu): small unit of mass used to measure atoms, subatomic particles, etc.

(f) Nucleus: small, dense, positively charged core of an atom containing protons and neutrons.

(g) Spectroscopy: the study of spectra produced by the excitation of atoms, especially experimental observation of optical spectra used to determine a substance's identity.

(h) Ground state: the state of lowest energy for an atom.

(i) Excited state: any state of higher energy for an atom.

(j) Atomic mass: the weighted average of all the isotopes of an element.

(k) Quantum Leap: change of an electron from one energy state to another within an atom.

(l) Photon: small packets of electromagnetic energy

2. Compare and contrast electrons, protons, and neutrons in terms of mass, charge and location.

PARTICLE	SYMBOL	CHARGE	MASS (amu)	LOCATION
electron	e^-	-1	≈0	orbit nucleus
proton	p^+	+1	≈1	inside nucleus
neutron	n^0	0	≈1	inside nucleus

3. What are the units used to express (a) wavelength, (b) frequency.
(a) any units of length (b) cycles/second or hertz.

4. Compare/contrast x-rays and infrared radiation in terms of their speed, wavelength, frequency and energy.

Both move at the speed of light however x-rays have shorter wavelengths which give them higher frequencies and energies than infrared radiation.

Chapter 4 Chemistry Study Guide Answers

Iris Stovall,Donald J. DeCoste

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Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble, 1994 *General Chemistry Study Guide Sixth Edition* Darrell D. Ebbing, 1999

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mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

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