

Electrons in Atoms

Section 5.1 Light and Quantized Energy

pages 144–145

Practice Problems

pages 140–143

1. Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of 4.90×10^{-7} m. What is the frequency of the light?

$$c = \lambda\nu$$

$$3.00 \times 10^8 \text{ m/s} = (4.90 \times 10^{-7} \text{ m})\nu$$

$$\nu = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}} = 6.12 \times 10^{14} \text{ Hz}$$

2. X rays can penetrate body tissues and are widely used to diagnose and treat disorders of internal body structures. What is the frequency of an X ray with a wavelength of 1.15×10^{-10} m?

$$c = \lambda\nu$$

$$3.00 \times 10^8 \text{ m/s} = (1.15 \times 10^{-10} \text{ m})\nu$$

$$\nu = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}} = 2.61 \times 10^{18} \text{ Hz}$$

3. After careful analysis, an electromagnetic wave is found to have a frequency of 7.8×10^6 Hz. What is the speed of the wave?

$$3.00 \times 10^8 \text{ m/s}$$

4. **Challenge** While an FM radio station broadcasts at a frequency of 94.7 MHz, an AM station broadcasts at a frequency of 820 kHz. What are the wavelengths of the two broadcasts? Which of the drawings below correspond to the FM station? To the AM station?



FM station:

$$c = \lambda\nu$$

$$\lambda = c/\nu$$

$$\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{94.7 \times 10^6 \text{ s}^{-1}} = 3.17 \text{ m}$$

FM: drawing b

AM station:

$$c = \lambda\nu$$

$$\lambda = c/\nu$$

$$\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{8.20 \times 10^5 \text{ s}^{-1}} = 366 \text{ m}$$

The 820 kHz carrier wave has the greater wavelength.

AM: drawing a

5. Calculate the energy possessed by a single photon of each of the following types of electromagnetic radiation.

a. $6.32 \times 10^{20} \text{ s}^{-1}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(6.32 \times 10^{20} \text{ s}^{-1})$$

$$E_{\text{photon}} = 4.19 \times 10^{-13} \text{ J}$$

b. $9.50 \times 10^{13} \text{ Hz}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(9.50 \times 10^{13} \text{ s}^{-1})$$

$$E_{\text{photon}} = 6.29 \times 10^{-20} \text{ J}$$

c. $1.05 \times 10^{16} \text{ s}^{-1}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(1.05 \times 10^{16} \text{ s}^{-1})$$

$$E_{\text{photon}} = 6.96 \times 10^{-18} \text{ J}$$

Chapter 5 Electrons In Atoms Solution Manual

Fujia Yang,Joseph H Hamilton

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