

Electrons in Atoms

Section 5.1 Light and Quantized Energy

pages 134–145

Practice Problems

pages 140–143

- Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of 4.90×10^{-7} m. What is the frequency of the light?

$$c = \lambda \nu$$

$$3.00 \times 10^8 \text{ m/s} = (4.90 \times 10^{-7} \text{ m})\nu$$

$$\nu = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}} = 6.12 \times 10^{14} \text{ Hz}$$

- X rays can penetrate body tissues and are widely used to diagnose and treat disorders of internal body structures. What is the frequency of an X ray with a wavelength of 1.15×10^{-10} m?

$$c = \lambda \nu$$

$$3.00 \times 10^8 \text{ m/s} = (1.15 \times 10^{-10} \text{ m})\nu$$

$$\nu = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}} = 2.61 \times 10^{18} \text{ Hz}$$

- After careful analysis, an electromagnetic wave is found to have a frequency of 7.8×10^6 Hz. What is the speed of the wave?

$$3.00 \times 10^8 \text{ m/s}$$

- Challenge** While an FM radio station broadcasts at a frequency of 94.7 MHz, an AM station broadcasts at a frequency of 820 kHz. What are the wavelengths of the two broadcasts? Which of the drawings below correspond to the FM station? To the AM station?



FM station:

$$c = \lambda \nu$$

$$\lambda = c / \nu$$

$$\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{9.47 \times 10^7 \text{ s}^{-1}} = 3.17 \text{ m}$$

FM: drawing b

AM station:

$$c = \lambda \nu$$

$$\lambda = c / \nu$$

$$\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{8.20 \times 10^5 \text{ s}^{-1}} = 366 \text{ m}$$

The 820 kHz carrier wave has the greater wavelength.

AM: drawing a

- Calculate the energy possessed by a single photon of each of the following types of electromagnetic radiation.

a. $6.32 \times 10^{20} \text{ s}^{-1}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(6.32 \times 10^{20} \text{ s}^{-1})$$

$$E_{\text{photon}} = 4.19 \times 10^{-13} \text{ J}$$

b. $9.50 \times 10^{13} \text{ Hz}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(9.50 \times 10^{13} \text{ s}^{-1})$$

$$E_{\text{photon}} = 6.29 \times 10^{-20} \text{ J}$$

c. $1.05 \times 10^{16} \text{ s}^{-1}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(1.05 \times 10^{16} \text{ s}^{-1})$$

$$E_{\text{photon}} = 6.96 \times 10^{-18} \text{ J}$$

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Fujia Yang, Joseph H Hamilton



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Table of Contents Chapter 5 Electrons In Atoms Solution Manual

1. Understanding the eBook Chapter 5 Electrons In Atoms Solution Manual
 - The Rise of Digital Reading Chapter 5 Electrons In Atoms Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 5 Electrons In Atoms Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 5 Electrons In Atoms Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 5 Electrons In Atoms Solution Manual
 - Personalized Recommendations
 - Chapter 5 Electrons In Atoms Solution Manual User Reviews and Ratings
 - Chapter 5 Electrons In Atoms Solution Manual and Bestseller Lists
5. Accessing Chapter 5 Electrons In Atoms Solution Manual Free and Paid eBooks
 - Chapter 5 Electrons In Atoms Solution Manual Public Domain eBooks
 - Chapter 5 Electrons In Atoms Solution Manual eBook Subscription Services
 - Chapter 5 Electrons In Atoms Solution Manual Budget-Friendly Options
6. Navigating Chapter 5 Electrons In Atoms Solution Manual eBook Formats

- ePub, PDF, MOBI, and More
- Chapter 5 Electrons In Atoms Solution Manual Compatibility with Devices
- Chapter 5 Electrons In Atoms Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 5 Electrons In Atoms Solution Manual
 - Highlighting and Note-Taking Chapter 5 Electrons In Atoms Solution Manual
 - Interactive Elements Chapter 5 Electrons In Atoms Solution Manual
- 8. Staying Engaged with Chapter 5 Electrons In Atoms Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 5 Electrons In Atoms Solution Manual
- 9. Balancing eBooks and Physical Books Chapter 5 Electrons In Atoms Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 5 Electrons In Atoms Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 5 Electrons In Atoms Solution Manual
 - Setting Reading Goals Chapter 5 Electrons In Atoms Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 5 Electrons In Atoms Solution Manual
 - Fact-Checking eBook Content of Chapter 5 Electrons In Atoms Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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