

## CHAPTER 9 REVIEW

***Stoichiometry*****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ The coefficients in a chemical equation represent the
  - (a) masses in grams of all reactants and products.
  - (b) relative number of moles of reactants and products.
  - (c) number of atoms of each element in each compound in a reaction.
  - (d) number of valence electrons involved in a reaction.
2. \_\_\_\_\_ Which of the following would not be studied within the topic of stoichiometry?
  - (a) the mole ratio of Al to Cl in the compound aluminum chloride
  - (b) the mass of carbon produced when a known mass of sucrose decomposes
  - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
  - (d) the amount of energy required to break the ionic bonds in  $\text{CaF}_2$
3. \_\_\_\_\_ A balanced chemical equation allows you to determine the
  - (a) mole ratio of any two substances in the reaction.
  - (b) energy released in the reaction.
  - (c) electron configuration of all elements in the reaction.
  - (d) reaction mechanism involved in the reaction.
4. \_\_\_\_\_ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
  - (a) reaction sequence.
  - (b) bond energy.
  - (c) mole ratio.
  - (d) element proportion.
5. Given the reaction represented by the following unbalanced equation:  
 $\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$ 
  - a. Balance the equation.

- |       |  |
|-------|--|
| _____ | b. What is the mole ratio of $\text{NO}_2$ to $\text{O}_2$ ?   |
| _____ | c. If 20.0 mol of $\text{NO}_2$ form, how many moles of $\text{O}_2$ must have been consumed?              |
| _____ | d. Twice as many moles of $\text{NO}_2$ form as moles of $\text{N}_2\text{O}$ are consumed. True or False? |
| _____ | e. Twice as many grams of $\text{NO}_2$ form as grams of $\text{N}_2\text{O}$ are consumed. True or False? |

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