

## 12

## STOICHIOMETRY

## Practice Problems

In your notebook, solve the following problems.

## SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

11. A car dealer has 100 bags of apples, 20 crates of peaches, and 10 tubs (pounds) of cherries. Write a balanced equation that fits this situation. How many apples are needed to make one peach?
12. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate,  $2\text{KClO}_3(s)$ . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
13. Using the equation from problem 12, how many moles of oxygen are produced from 1.0 moles of potassium chlorate?
14. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from  $2.0 \times 10^{24}$  molecules of oxygen? How many moles of water are produced from  $22.0 \times 10^{24}$  molecules of oxygen?

## SECTION 12.2 CHEMICAL CALCULATIONS

15. Calculate the number of moles of hydrogen chloride produced from 30 moles of hydrochloric acid.



16. Calculate the number of moles of chlorine needed to form 1.0 moles of iron(III) chloride.



17. Calculate the number of grams of nitrogen dioxide that are produced from 4.0 moles of nitric oxide.



18. Calculate the mass, in grams, produced from the decomposition of 75.0 g of potassium chlorate.



19. Calculate the mass, in grams, needed to react with chlorine to produce 164 g of silver chloride. (Refer to Table 1 for the molar mass of each element.)

20. How many liters of carbon monoxide at  $500^\circ\text{C}$  are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



21. Calculate the number of moles of oxygen gas needed to produce 4.00 liters of nitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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### **Table of Contents Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers**

1. Understanding the eBook Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - The Rise of Digital Reading Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Personalized Recommendations
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers User Reviews and Ratings
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers and Bestseller Lists

5. Accessing Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers Free and Paid eBooks
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers Public Domain eBooks
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers eBook Subscription Services
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers Budget-Friendly Options
6. Navigating Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers Compatibility with Devices
  - Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Highlighting and Note-Taking Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Interactive Elements Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
8. Staying Engaged with Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
9. Balancing eBooks and Physical Books Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Setting Reading Goals Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Fact-Checking eBook Content of Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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