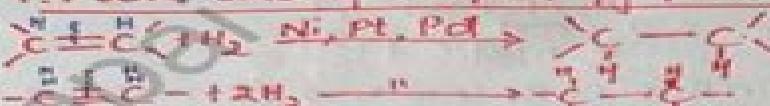


HYDROCARBON

MOP of Alkane

1. From unsaturated hydrocarbons



DM → π bond ke 4-4 se foot qd.

Exothermic ROR: Alkyne > Alkene

4 MCts formed ROR $\propto \frac{1}{S \cdot H}$

Syn addition \propto stability

Add of conj. diene $\text{CH}_2=\text{CH}-\text{CH}=\text{CH}_2 \rightarrow \text{C}_6\text{H}_6$

RuNi → alloy of Ni & Al in NaOH

Galvanic cell better than Ni (at 100°C)

If RuNi used → Sabatier's senderson's

& $[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$ used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	Ethyne
Trans	Syn	Thene
Cis	Anti	Thene
Trans	Anti	Ethyne

2. From alkyl halide



X halao \propto leaving

ROR: $\text{R}-\text{Cl} < \text{R}-\text{Br} < \text{R}-\text{I}$ ROR $\propto \frac{1}{S \cdot E}$

3. Wurtz reaction



produce higher alkane with even no. of C's

R-X → coupling

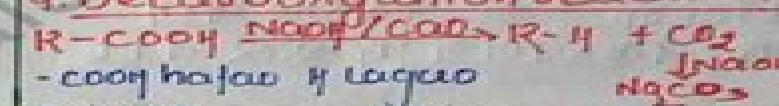
(1, 2, 3)
R-X → disproportionation

intra molecular Wurtz faster than inter molecular but prod should be stable i.e. A-A nahi banana hai.

ROR $\frac{1}{S \cdot H}$ RRX > 2° RX > 3° RX

for same R: RF < RCl < RBr < RI

4. Decarboxylation reaction



-COO₂ halao \propto leaving

ROR \propto stability of intermediate (R⁻)

5. Kolbe's electrolysis

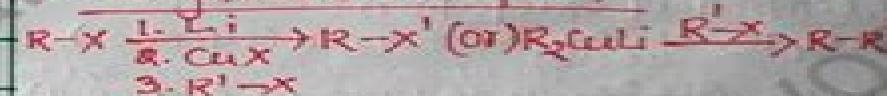


Addition Reaction

6. Frankland reaction



7. Cosyhouse reaction



Attack of R⁻ → RDS

ROR $\propto \frac{1}{S \cdot H}$ of R'-X

If same R:



BE \downarrow → leaving tendency \uparrow

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