



CHEMISTRY

CHEAT SHEET

GAS LAWS



GAS LAWS CHEAT SHEET

Boyle's Law states that the pressure of a gas is inversely proportional to its volume, provided the temperature and the amount of gas remain constant. This relationship can be expressed mathematically as:

$$P_1 V_1 = P_2 V_2$$

Variable	Unit	Symbol
Pressure	atm, mmHg, kPa	P
Volume	L, mL, m³	V
Temperature	°C, K	T
Amount of gas	mol	n

Boyle's Law

Boyle's Law states that the pressure of a gas is inversely proportional to its volume, provided the temperature and the amount of gas remain constant.

$$P_1 V_1 = P_2 V_2$$

Boyle's Law can be expressed graphically as a plot of Pressure (P) versus Volume (V), which shows a hyperbolic curve.



Charles' Law

Charles' Law states that the volume of a gas is directly proportional to its temperature, provided the pressure and the amount of gas remain constant.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Charles' Law can be expressed graphically as a plot of Volume (V) versus Temperature (T), which shows a straight line passing through the origin.



Avogadro's Law

Avogadro's Law states that the volume of a gas is directly proportional to the number of moles of gas, provided the pressure and temperature remain constant.

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

Avogadro's Law can be expressed graphically as a plot of Volume (V) versus Number of moles (n), which shows a straight line passing through the origin.



Ideal Gas Law

The Ideal Gas Law is a combination of Boyle's, Charles', and Avogadro's Laws. It can be expressed mathematically as:

$$PV = nRT$$

Where: P = pressure (atm), V = volume (L), n = number of moles, R = gas constant (0.0821 L·atm/mol·K), and T = temperature (K).

Boyle's Law is a special case of the Ideal Gas Law where temperature and the number of moles are constant.

$$P_1 V_1 = P_2 V_2$$

$$P_1 = \frac{nRT}{V_1}$$

$$V_1 = \frac{nRT}{P_1}$$

$$P_2 = \frac{nRT}{V_2}$$

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M Lipman



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