

Name \_\_\_\_\_ answer key \_\_\_\_\_ Date \_\_\_\_\_

# Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl<sub>3</sub>.

9 FeCl <sub>3</sub>	1 mole FeCl <sub>3</sub>	mole NaCl
9 FeCl <sub>3</sub>	9 FeCl <sub>3</sub>	mole FeCl

Calculate the amount of Fe<sub>2</sub> (SO<sub>4</sub>)<sub>3</sub> in grams produce 215 g Na<sub>2</sub>SO<sub>4</sub>.

9 Na <sub>2</sub> SO <sub>4</sub>	1 mole Na <sub>2</sub> SO <sub>4</sub>	mole Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
9 Na <sub>2</sub> SO <sub>4</sub>	9 Na <sub>2</sub> SO <sub>4</sub>	mole Na <sub>2</sub> SO <sub>4</sub>	1 mole Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>

Calculate the amount in grams of FeCl<sub>3</sub> produced with 95g of Fe (SO<sub>4</sub>)<sub>3</sub>

9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	1 mole Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	mole FeCl <sub>3</sub>	9 FeCl <sub>3</sub>
9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	9 Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	mole Fe(SO <sub>4</sub> ) <sub>3</sub>	1 mole FeCl <sub>3</sub>

Calculate the amount of moles of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> needed to react with 123.5 g of NaCl

9 NaCl	1 mole NaCl	mole Fe(SO <sub>4</sub> ) <sub>3</sub>
9 NaCl	9 NaCl	mole NaCl

Based on the following chemical equation, please answer the following questions!



9 Ca(HCO <sub>3</sub> ) <sub>2</sub>	1 mole Ca(HCO <sub>3</sub> ) <sub>2</sub>	mole CaCO <sub>3</sub>
9 Ca(HCO <sub>3</sub> ) <sub>2</sub>	9 Ca(HCO <sub>3</sub> ) <sub>2</sub>	mole Ca(HCO <sub>3</sub> ) <sub>2</sub>

# Chemistry Study Guide The Mole Answers

**John William Hill, Dorothy M. Feigl**



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### **Table of Contents Chemistry Study Guide The Mole Answers**

1. Understanding the eBook Chemistry Study Guide The Mole Answers
  - The Rise of Digital Reading Chemistry Study Guide The Mole Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide The Mole Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Study Guide The Mole Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide The Mole Answers
  - Personalized Recommendations
  - Chemistry Study Guide The Mole Answers User Reviews and Ratings
  - Chemistry Study Guide The Mole Answers and Bestseller Lists

5. Accessing Chemistry Study Guide The Mole Answers Free and Paid eBooks
  - Chemistry Study Guide The Mole Answers Public Domain eBooks
  - Chemistry Study Guide The Mole Answers eBook Subscription Services
  - Chemistry Study Guide The Mole Answers Budget-Friendly Options
6. Navigating Chemistry Study Guide The Mole Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Study Guide The Mole Answers Compatibility with Devices
  - Chemistry Study Guide The Mole Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Study Guide The Mole Answers
  - Highlighting and Note-Taking Chemistry Study Guide The Mole Answers
  - Interactive Elements Chemistry Study Guide The Mole Answers
8. Staying Engaged with Chemistry Study Guide The Mole Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Study Guide The Mole Answers
9. Balancing eBooks and Physical Books Chemistry Study Guide The Mole Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Study Guide The Mole Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide The Mole Answers
  - Setting Reading Goals Chemistry Study Guide The Mole Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide The Mole Answers
  - Fact-Checking eBook Content of Chemistry Study Guide The Mole Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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