

QUIZZZ

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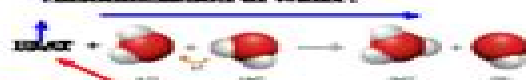
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Unit 8

Acids & Bases

8.1 Introduction to Acids & Bases

Autoionization of Water:



This is an endothermic process- increasing the heat will shift the reaction to produce more H_3O^+ and OH^- . K_w increases (more products), both pH and pOH will decrease but pure H₂O will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$$\text{pH} = 3.45$$

is sig. figs

In case of acidic solution $> [\text{H}^+] > [\text{OH}^-]$

$$\text{pH} < 7 \text{ at } 25^\circ\text{C}$$

In case of neutral solution $> [\text{H}^+] = [\text{OH}^-]$

$$\text{pH} = 7 \text{ at } 25^\circ\text{C}$$

In case of basic solution $> [\text{H}^+] < [\text{OH}^-]$

$$\text{pH} > 7 \text{ at } 25^\circ\text{C}$$

ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also known:

$$[\text{H}^+] = 10^{-\text{pH}} \text{ and } [\text{OH}^-] = 10^{-\text{pOH}}$$

3. The value of K_w for water at 0°C is 1×10^{-15} . What is the pOH of water at 0°C ?

a) 7.5

b) 8.0

c) 7.0

d) 6.5

Chemistry Unit 8 Review Answers

Marco Cascella

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Princeton Review AP Chemistry Premium Prep, 27th Edition The Princeton Review, 2025-08-05 PREMIUM PRACTICE FOR A PERFECT 5 WITH THE MOST PRACTICE ON THE MARKET Equip yourself to ace the newly digital AP Chemistry Exam with this comprehensive study guide Includes 7 full length practice tests with answer explanations timed online practice and thorough content reviews Techniques That Actually Work Tried and true strategies to help you avoid traps and beat the test Tips for pacing yourself and guessing logically Essential tactics to help you work smarter not harder Everything You Need for a High Score Updated to address the new digital exam Comprehensive content review for all test topics Online digital flashcards to review core content Study plans a handy list of key terms and concepts and more via your online Student Tools Premium Practice for AP Excellence 7 full length practice tests 5 in the book 2 online with detailed answer explanations Online tests provided as both digital versions with timer option to simulate exam experience online and as downloadable PDFs with interactive elements mimicking the exam interface Review of important laboratory procedures and equipment Practice drills at the end of each content review chapter plus step by step walk throughs of sample exam questions

CliffsNotes HESI A2 Science Cram Plan Michael Reid, 2021-04-13 A study guide for the HESI A2 science nursing school test that calendarizes a study plan for test takers depending on how much time they have left before taking the test Get a plan and make the most of the time you have left Whether you have two months one month or one week left before the exam you can turn to the experts at CliffsNotes for a trusted and achievable cram plan to ace the HESI A2 Science without ever breaking a sweat First you'll determine exactly how much time you have left to prepare for the exam Then you'll turn to the two month one month or one week cram plan for week by week and day by day schedules of the best way to focus your study according to your unique timeline Each stand alone plan includes Diagnostic tests help you pinpoint your strengths and weaknesses so you can focus your review on the topics in which you need the most help Subject areas review of material you should know for the exam biology chemistry anatomy and physiology and physics Practice exams with answers and detailed explanations

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Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow,1978

Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001 **Chemistry and Industry Review** ,1928 Chemistry Royal B. Leach,Galen Wood Ewing,1966 Chemistry: Matter & Change, Solving Problems - A

Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change

Qualitative Analysis and Ionic Equilibrium George H. Schenk, 1990 A supplement for courses with a qualitative analysis component this lab manual contains explanations of the chemistry of metal ions and anions It includes pre lab exercises experiments and lab reports *Pennsylvania Township News*, 1974 **McGraw-Hill Education 500 Review Questions**

for the MCAT: General Chemistry Mary Millhollon, Richard H. Langley, 2015-07-10 500 ways to pass the General Chemistry section of the new MCAT Intensive practice detailed answer explanations the best way to sharpen skills and prepare for the exam In anticipation of the fully revised 2015 MCAT 500 Review Questions for the MCAT General Chemistry has been updated to comprehensively cover the chemistry portion of the Chemical and Physical Foundations of Biological Systems section This book provides the problem solving practice you need to take the exam with confidence 500 questions organized by subject Follows the new MCAT format Complete explanations to every question given in the answer key

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