

**Matching**

Match each item with the correct statement below.

- |                  |                               |
|------------------|-------------------------------|
| a. Boyle's law   | d. Graham's law               |
| b. Charles's law | e. Gay-Lussac's law           |
| c. Dalton's law  | f. ideal gas law (mark A & B) |

- \_\_\_\_\_ 1) For a given mass of gas at constant temperature, the volume of the gas varies inversely with pressure.
- \_\_\_\_\_ 2) The volume of a fixed mass of gas is directly proportional to its Kelvin temperature, if the pressure is kept constant.
- \_\_\_\_\_ 3) The pressure of a gas is directly proportional to its Kelvin temperature if the volume is kept constant.
- \_\_\_\_\_ 4)  $P \cdot V = n \cdot R \cdot T$
- \_\_\_\_\_ 5) At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases.
- \_\_\_\_\_ 6) The rate at which a gas will effuse is inversely proportional to the square root of the gas's molar mass.

Match each item with the correct statement below.

- |                    |                     |
|--------------------|---------------------|
| a. effusion        | c. diffusion        |
| b. compressibility | d. partial pressure |

- \_\_\_\_\_ 7) a measure of how much the volume of matter decreases under pressure
- \_\_\_\_\_ 8) the pressure exerted by a gas in a mixture
- \_\_\_\_\_ 9) the escape of gas through a small hole in a container
- \_\_\_\_\_ 10) tendency of molecules to move to regions of lower concentration

**Multiple Choice**

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 11) Why is a gas easier to compress than a liquid or a solid?
- Its volume increases more under pressure than an equal volume of liquid does.
  - Its volume increases more under pressure than an equal volume of solid does.
  - The space between gas particles is much less than the space between liquid or solid particles.
  - The volume of a gas's particles is small compared to the overall volume of the gas.
- \_\_\_\_\_ 12) Why does the pressure inside a container of gas increase if more gas is added to the container?
- There is an increase in the number of collisions between particles and the walls of the container.
  - There is an increase in the temperature of the gas.
  - There is a decrease in the volume of the gas.
  - There is an increase in the force of the collisions between the particles and the walls of the container.
- \_\_\_\_\_ 13) How does the gas propellant move when an aerosol can is used?
- from a region of high pressure to a region of lower pressure
  - from a region of high pressure to a region of equally high pressure
  - from a region of low pressure to a region of higher pressure

# Chapter 14 Gases Study Guide Answer Key

**Emilie Sanchez**



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