

Acids and Bases Summary

Properties of Aqueous Acids:

1. Acids are electrolytes: aqueous solutions of acids ionize when added to water and conduct electric current.
2. Traditional Acids contain hydrogen and react with active metals to produce hydrogen gas (H_2) by single replacement reactions.
3. Acids change the color of dyes known as acid base indicators; i.e. acids turn blue litmus red, turns bromothymol blue to yellow, and leaves phenolphthalein clear.
4. Neutralizes bases: acids react with bases to produce salts and water. Equivalent amounts of acid and base are neutralized (neutralization reactions) to produce water and a salt (ionic compound).
$$NaOH_{(aq)} + HCl_{(aq)} \rightarrow NaCl_{(aq)} + H_2O_{(aq)}$$
5. Sour Taste - Weak acids taste sour, i.e.: citrus fruits (lemons, oranges, grapefruit) contain citric acid; apples - malic acid; vinegar - acetic acid; sour milk - lactic acid; wine - tartaric acid; carbonated beverages - phosphoric acid.
6. Acids have a pH less than 7.0
7. Strong Acids: Acids that are completely or nearly completely ionized (dissolved) including HCl , H_2SO_4 , HBr , HNO_3 , etc. Strong acids in concentrated solutions are very corrosive.
Weak Acids: those acids that are only slightly ionized (dissolved) including CH_3COOH , H_3PO_4 , HF , etc.

Properties of Aqueous Bases:

1. Bases are electrolytes: aqueous solutions of bases conduct electricity.
2. Bases, like acids, change the color of acid base indicators: i.e. bases turn red litmus blue, leaves bromothymol blue, blue, and turns phenolphthalein pink.
3. Bases neutralize acids to produce salts and water. $NaOH_{(aq)} + HCl_{(aq)} \rightarrow NaCl_{(aq)} + H_2O_{(aq)}$
4. Bases have a bitter taste.
5. Dilute aqueous solutions of bases feel slippery to the skin (soap)
6. Bases have a pH greater than 7.0
7. Strong Bases: those bases that are completely or nearly completely ionized including KOH , $NaOH$, $Ba(OH)_2$, $Sr(OH)_2$, and $Ca(OH)_2$. All active metal hydroxides (group 1 and 2 metals) form strong bases. Strong bases like strong acids are very corrosive.
Weak Bases: those bases that are only slightly ionized including NH_4OH , and all others.

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Ian Pickup



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Making it tangible. Learning outcomes in science education Sascha Bernholt, Knut Neumann, Peter Nentwig, 2012 One of the central features in current educational reforms is a focus on learning outcomes Many countries have established or revised standards to describe what teachers are supposed to teach and students are expected to learn More recently the emphasis has shifted to considerations of how standards can be operationalized in order to make the outcomes of educational efforts more tangible This book is the result of a symposium held in Kiel that was arranged by two science education groups one at the IPN Leibniz Institute for Science and Mathematics Education at the University of Kiel in Germany and the other at the University of York UK The seminar brought together renowned experts from 12 countries with different notions of the nature and quality of learning outcomes The aim was to clarify central conceptions and approaches for a better understanding among the international science education community The book is divided into five parts In Part A the organizers set the scene describing the rationale for arranging the symposium Part B provides a broad overview about different approaches challenges and pitfalls on the road to the clarification of meaningful and fruitful learning outcomes The set of papers in Part

C provides deep insights into different although comparable approaches which aim to frame to assess and to promote learning and learning outcomes in science education Smaller projects are presented as well as broad coordinated national programs The papers in Part D outline the individual historical development from different national perspectives reflecting the deficits and problems that led to current reforms Finally a summary of the organizers analyses the conclusions from different vantage points **Basic Nursing** Patricia Ann Potter, Anne Griffin Perry, 1991 Approach to Chemistry ,1962

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