

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY**Covalent Bonding****Section 11 The Covalent Bond**

In your textbook, read about the nature of covalent bonds.

Underline the terms below just now to complete the passage.

covalent bond	valence	valence	attractive	pull toward
When sharing of electrons occurs, the attraction between atoms that results is called	(a) covalent bond	When each atom has a limited total electron capacity, and the product is (b) neutrophilic . When two or more atoms bond by means of sharing during the resulting patch is called an		

(c) **molophile**. If the electrons shared are attracted between the two atoms, the attraction is called an (d) **sigma bond**. If the electrons shared are attracted between the two atoms, the repulsion of electrons is called an (e) **pi bond**.

In your textbook, read about single and double bonds, and bond strength.

Circle the letter of the theory that best completes the statement in answer to the question.

4. In what form do electrons in two hydrogen atoms, and oxygen atom(s) move?

a. as single atoms
(f) a molecule containing two atoms
b. as molecules containing two atoms

5. How many electrons are shared in a double covalent bond?

a. one
(g) two
b. one

6. Bond length is the distance between

a. two molecules of the same substance
b. the electrons in two different atoms
(h) the nuclei of two different atoms

7. What is the following relationship referring to bond length in generally untrue?

(i) As atoms get larger, the length of the bond
b. the shorter the bond, the weaker the bond
c. the shorter the bond, the lower the electrons in it
d. the shorter the bond, the larger the bond length

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Section 12 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

(a) _____ 1. Every molecular compound is generally composed of a metal and a nonmetal.

(b) _____ 2. The second element in the formula of a binary compound is named using the suffix -ide.

(c) _____ 3. The prefix two is added in front of atoms.

(d) _____ 4. The prefix two is added in front of atoms.

(e) _____ 5. In naming the first element in a formula, the prefix mono is not used.

(f) _____ 6. For binary acids, the hydrogen part of the compound is named using the prefix hydro-.

(g) _____ 7. An oxygen atom is only two electrons.

(h) _____ 8. If the name of the element in a binary acid is on the end of the name, remove the suffix -ide.

In your textbook, read about naming molecular compounds and acids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

1.	Li	1. hydride acid
2.	Na	2. oxygen molecule
3.	O ₂	3. carbon monoxide
4.	CO	4. carbon dioxide
5.	CO ₂	5. nitric acid
6.	NO ₂	6. sulfur acid
7.	NO	7. sulfuric acid
8.	SO ₂	8. hydrochloric acid

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Section 13 Chemistry and Change • Chapter 1

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reader will excuse me if i conceal the date or any other fact by which he might trace the actual occurrence we had been out for one of our evening rambles holmes and i and had returned about six o clock on a cold frosty winter
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