

balancing Equations

Name: _____

Date: _____

Page _____

Introduction/Questions

- Write the equation for the reaction of sodium and chlorine (chlorine from sodium chloride).



- Write the equation for the reaction of sodium nitrate and sodium chloride to produce sodium chloride and sodium nitrate.



Information: Subscripts and Coefficients

It is important to understand that atoms are never really "destroyed" or "created". For example, in CaCl_2 there are 1 calcium atom and 2 chlorine atoms. In CaCO_3 there are 1 calcium atom, 1 carbon atom, and 3 oxygen atoms.

A coefficient tells you how many molecules (or atoms) there are. For example, in 2NaCl the "2" is the coefficient. The "Na" tells us that there are three molecules of sodium present. In this equation, " 2NaCl " there are a total of 2 sodium atoms and 2 chlorine atoms.

Balance Chemical Equations

- Balance oxygen atoms in H_2SO_4 . 3/1
- Balance oxygen atoms in the reaction $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_3 + \text{H}_2\text{O}_2$



Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride (halogenation):

$$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$$

Remember that when chlorine is in Cl_2 it is always sodium is Cl_2 . On the product side of the equation, there are only one sodium atom and one chlorine atom. To balance the equation, we need to add a coefficient of 2 to the product side. This gives us:

$$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$$

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