

The Mole Study Guide

| Topic | Definition | Example |
|--|--|--|
| Definition of a Mole | Avogadro's number of particles | 6.02 x 10 ²³ |
| Converting between moles and particles | $\frac{\text{moles}}{\text{Avogadro's number}} = \text{particles}$ | 1 mole of H_2O = 6.02 x 10 ²³ molecules |
| Converting between moles and mass | $\frac{\text{moles}}{\text{molar mass}} = \text{mass}$ | 1 mole of H_2O = 18.0 g |
| Converting between mass and moles | $\frac{\text{mass}}{\text{molar mass}} = \text{moles}$ | 18.0 g of H_2O = 1 mole |
| Converting between moles and volume | $\frac{\text{moles}}{\text{molar volume}} = \text{volume}$ | 1 mole of H_2O = 18.0 L |
| Converting between volume and moles | $\frac{\text{volume}}{\text{molar volume}} = \text{moles}$ | 18.0 L of H_2O = 1 mole |
| Converting between moles and energy | $\frac{\text{moles}}{\text{molar energy}} = \text{energy}$ | 1 mole of H_2O = 18.0 kJ |
| Converting between energy and moles | $\frac{\text{energy}}{\text{molar energy}} = \text{moles}$ | 18.0 kJ of H_2O = 1 mole |
| Converting between moles and time | $\frac{\text{moles}}{\text{molar time}} = \text{time}$ | 1 mole of H_2O = 18.0 s |
| Converting between time and moles | $\frac{\text{time}}{\text{molar time}} = \text{moles}$ | 18.0 s of H_2O = 1 mole |
| Converting between moles and space | $\frac{\text{moles}}{\text{molar space}} = \text{space}$ | 1 mole of H_2O = 18.0 m ³ |
| Converting between space and moles | $\frac{\text{space}}{\text{molar space}} = \text{moles}$ | 18.0 m ³ of H_2O = 1 mole |
| Converting between moles and mass | $\frac{\text{moles}}{\text{molar mass}} = \text{mass}$ | 1 mole of H_2O = 18.0 g |
| Converting between mass and moles | $\frac{\text{mass}}{\text{molar mass}} = \text{moles}$ | 18.0 g of H_2O = 1 mole |
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Chemistry The Mole Study Guide

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