

# balancing Equations

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## Introduction/Questions

- Write the equation for the reaction of sodium and chlorine (chlorine gas) to form sodium chloride.



- Write the equation for the reaction of sodium nitrate and sodium chloride to produce sodium chloride and sodium nitrate.



## Information: Subscripts and Coefficients

It is important to understand that when you have subscripts in a compound, the number of Cl<sub>2</sub> molecules is the subscript and the Cl<sub>2</sub> is that there are two chlorine atoms bonded to one carbon.

A coefficient tells you how many molecules of a compound there are, that is a different way. For example in the equation: "1 NaCl" the NaCl is the coefficient. The three tells us that there are three molecules of sodium present. In the equation "1 NaCl" there are a total of 1 sodium atoms and 1 chlorine atom.

## Balancing Chemical Equations

- How many oxygen atoms are present in 4Na<sub>2</sub>O<sub>2</sub>? 8/2
- How many oxygen atoms are in each of the following:



## Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride. This equation is:

$$Na + Cl_2 \rightarrow NaCl$$

Remember that when chlorine is in Cl<sub>2</sub> it is always written as Cl<sub>2</sub>. On the product side of the equation, there are only one sodium atom and one chlorine atom. On the reactant side, there are only one sodium atom and two chlorine atoms. In order for the equation to make sense, we need to balance the equation. This can be done by adding a "2" to the product side:



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