

## 16 • Chemical Equilibrium

1. Consider the reaction system.  
 $\text{CoO}(s) + \text{H}_2(g) \rightleftharpoons \text{Co}(s) + \text{H}_2\text{O}(g)$ .

The equilibrium constant expression is

a)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$       d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$   
 b)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$       c)  $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$   
 (Handwritten note:  $\text{CoO} \rightarrow \text{Co}$ )  
 (Handwritten note:  $\text{H}_2 \rightarrow \text{H}_2\text{O}$ )

2. Given the equilibrium,  
 $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$ , if this equilibrium is established by beginning with equal number of moles of  $\text{SO}_2$  and  $\text{O}_2$  in a 1.0 Liter bulb, then the following *must* be true at equilibrium:  
 a)  $[\text{SO}_2] = [\text{SO}_3]$       d)  $[\text{SO}_2] < [\text{O}_2]$   
 (Handwritten note:  $2[\text{SO}_2] = 2[\text{SO}_3]$ ) e)  $[\text{SO}_2] > [\text{O}_2]$   
 c)  $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole  $\text{NO}$ , 0.200 mol  $\text{Cl}_2$  and 0.500 mol  $\text{ClNO}$  were placed in a 25.0 Liter container. The following equilibrium is established:  $2\text{ClNO}(g) \rightleftharpoons 2\text{NO}(g) + \text{Cl}_2(g)$

3. At equilibrium, 0.600 mol of  $\text{ClNO}$  was present. The number of *moles* of  $\text{Cl}_2$  present at equilibrium is  
 a) 0.050      (d) 0.200  
 b) 0.100      e) 0.250  
 c) 0.150

## PRACTICE TEST

4. The equilibrium constant,  $K_c$ , is:  
 a)  $4.45 \times 10^{-4}$       d) 0.167  
 b)  $6.67 \times 10^{-4}$       e) 1500  
 (Handwritten note:  $\text{c) } 0.111$ )

5. At  $985^\circ\text{C}$ , the equilibrium constant for the reaction,  
 $\text{H}_2(g) + \text{CO}_2(g) \rightleftharpoons \text{H}_2\text{O}(g) + \text{CO}(g)$

is 1.63. What is the equilibrium constant for the reverse reaction?  
 a) 1.63      d) 0.613  
 b) 0.815      (Handwritten note:  $\text{c) } 1.00$ )  
 c) 2.66

6. What is the relationship between  $K_p$  and  $K_c$  for the reaction,  $2\text{ICl}(g) \rightleftharpoons \text{I}_2(g) + \text{Cl}_2(g)$ ?  
 (Handwritten note:  $\text{a) } K_p = K_c(RT)^{-1}$ ) d)  $K_p = K_c$   
 b)  $K_p = K_c(RT)$       e)  $K_p = K_c(2RT)$   
 c)  $K_p = K_c(RT)^2$

7. For the reaction  $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ ,  $K_p$  at  $25^\circ\text{C}$  is 7.3, when all partial pressures are expressed in atmospheres. What is  $K_c$  for this reaction? [ $R=0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$ ]  
 (Handwritten note:  $\text{a) } 4270$ ) d) 179  
 b) 0.0119      e) 2.06  
 c) 0.291

# Ch 16 Chemical Equilibrium Problem Set 1

**David Elliott Goldberg**

## Ch 16 Chemical Equilibrium Problem Set 1:

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Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb Wade A. Freeman,1986

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1. Understanding the eBook Ch 16 Chemical Equilibrium Problem Set 1
  - The Rise of Digital Reading Ch 16 Chemical Equilibrium Problem Set 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Ch 16 Chemical Equilibrium Problem Set 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ch 16 Chemical Equilibrium Problem Set 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 16 Chemical Equilibrium Problem Set 1
  - Personalized Recommendations
  - Ch 16 Chemical Equilibrium Problem Set 1 User Reviews and Ratings
  - Ch 16 Chemical Equilibrium Problem Set 1 and Bestseller Lists
5. Accessing Ch 16 Chemical Equilibrium Problem Set 1 Free and Paid eBooks

- Ch 16 Chemical Equilibrium Problem Set 1 Public Domain eBooks
- Ch 16 Chemical Equilibrium Problem Set 1 eBook Subscription Services
- Ch 16 Chemical Equilibrium Problem Set 1 Budget-Friendly Options

6. Navigating Ch 16 Chemical Equilibrium Problem Set 1 eBook Formats

- ePUB, PDF, MOBI, and More
- Ch 16 Chemical Equilibrium Problem Set 1 Compatibility with Devices
- Ch 16 Chemical Equilibrium Problem Set 1 Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Ch 16 Chemical Equilibrium Problem Set 1
- Highlighting and Note-Taking Ch 16 Chemical Equilibrium Problem Set 1
- Interactive Elements Ch 16 Chemical Equilibrium Problem Set 1

8. Staying Engaged with Ch 16 Chemical Equilibrium Problem Set 1

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Ch 16 Chemical Equilibrium Problem Set 1

9. Balancing eBooks and Physical Books Ch 16 Chemical Equilibrium Problem Set 1

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Ch 16 Chemical Equilibrium Problem Set 1

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Ch 16 Chemical Equilibrium Problem Set 1

- Setting Reading Goals Ch 16 Chemical Equilibrium Problem Set 1
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Ch 16 Chemical Equilibrium Problem Set 1

- Fact-Checking eBook Content of Ch 16 Chemical Equilibrium Problem Set 1
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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