

CHAPTER 8 STUDY GUIDE

Covalent Bonding

Section 8.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond

molecule

sigma bond

exothermic

pi bond

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1)** _____. When such an attachment is formed, bond dissociation energy is released, and the process is **(2)** _____. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3)** _____. If the electrons shared are centered between the two atoms, the attachment is called a(n) **(4)** _____. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)** _____.

In your textbook, read about single and multiple bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

6. In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
 - a. as single atoms
 - b. as molecules containing two atoms
 - c. as molecules containing three atoms
 - d. as molecules containing four atoms
7. How many electrons are shared in a double covalent bond?
 - a. none
 - b. one
 - c. two
 - d. four
8. Bond length is the distance between
 - a. two molecules of the same substance.
 - b. the electrons in two attached atoms.
 - c. the nuclei of two attached atoms.
 - d. the orbitals of two attached atoms.
9. Which of the following relationships relating to bond length is generally correct?
 - a. the shorter the bond, the stronger the bond
 - b. the shorter the bond, the weaker the bond
 - c. the shorter the bond, the fewer the electrons in it
 - d. the shorter the bond, the lower the bond dissociation energy

Chemistry Study Guide Covalent Bonding

Paul B. Kelter



Chemistry Study Guide Covalent Bonding:

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Introductory Chemistry Study Guide Holder, 2003-05

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems

Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1992

Chemistry, Study Guide Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460

Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney William H. Myers, William T. Mooney, 1980

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