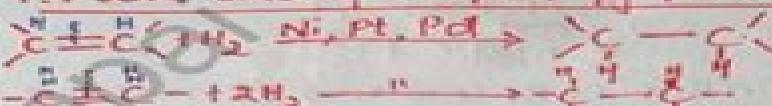


HYDROCARBON

MOP of Alkane

1. From unsaturated hydrocarbons



DM → π bond ke 4-4 se foot qd.

Exothermic ROR: Alkyne > Alkene

4 MCts formed ROR $\propto \frac{1}{S \cdot H}$

Syn addition α stability

Add of conj. diene $\text{CH}_2=\text{CH}-\text{CH}=\text{CH}_2 \rightarrow \text{C}_6\text{H}_6$

RuNi → alloy of Ni & Al in NaOH

Galvanic cell better than Ni (at 100°C)

If RuNi used → Sabatier's senderson's

& $[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$ used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	Ethyne
Trans	Syn	Thene
Cis	Anti	Thene
Trans	Anti	Ethyne

2. From alkyl halide



X halao \downarrow leaving

ROR: $\text{R}-\text{Cl} < \text{R}-\text{Br} < \text{R}-\text{I}$ ROR $\propto \frac{1}{S \cdot E}$

3. Wurtz reaction



produce higher alkane with even no. of C's

R-X → coupling

(1,2')

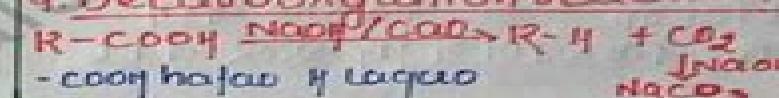
R-X → disproportionation

(3')
intra molecular Wurtz faster than inter molecular but prod should be stable i.e. A.A nahi banana hai.

ROR $\frac{1}{S \cdot H}$ R-R > 2' RX > 3' RX

If same R: RF < RCl < RBr < RI

4. Decarboxylation reaction



ROR \propto stability of intermediate (R⁻)

5. Kolbe's electrolysis



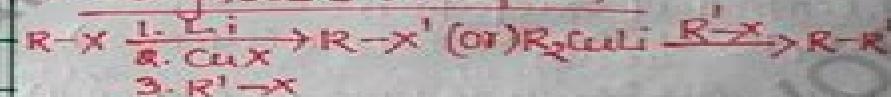
Cathode $\rightarrow \text{H}_2$ Anode \rightarrow Alkane + CO₂

Addition Reaction

6. Frankland reaction



7. Cosyhouse reaction



Attack of R⁻ → RDS

ROR $\propto \frac{1}{S \cdot H}$ of R'-X

If same R:



BE \downarrow → leaving tendency \uparrow

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Jill Meryl Levy

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Table of Contents Chemistry Study Guide Hydrocarbons Answer

1. Understanding the eBook Chemistry Study Guide Hydrocarbons Answer
 - The Rise of Digital Reading Chemistry Study Guide Hydrocarbons Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Hydrocarbons Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Hydrocarbons Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Hydrocarbons Answer
 - Personalized Recommendations

- Chemistry Study Guide Hydrocarbons Answer User Reviews and Ratings
- Chemistry Study Guide Hydrocarbons Answer and Bestseller Lists

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- Chemistry Study Guide Hydrocarbons Answer eBook Subscription Services
- Chemistry Study Guide Hydrocarbons Answer Budget-Friendly Options

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- ePUB, PDF, MOBI, and More
- Chemistry Study Guide Hydrocarbons Answer Compatibility with Devices
- Chemistry Study Guide Hydrocarbons Answer Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Study Guide Hydrocarbons Answer
- Highlighting and Note-Taking Chemistry Study Guide Hydrocarbons Answer
- Interactive Elements Chemistry Study Guide Hydrocarbons Answer

8. Staying Engaged with Chemistry Study Guide Hydrocarbons Answer

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Study Guide Hydrocarbons Answer

9. Balancing eBooks and Physical Books Chemistry Study Guide Hydrocarbons Answer

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Study Guide Hydrocarbons Answer

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Study Guide Hydrocarbons Answer

- Setting Reading Goals Chemistry Study Guide Hydrocarbons Answer
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Study Guide Hydrocarbons Answer

- Fact-Checking eBook Content of Chemistry Study Guide Hydrocarbons Answer

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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