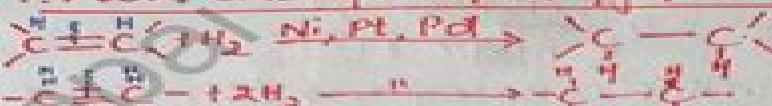


HYDROCARBON

MOP of Alkane

1. From unsaturated hydrocarbons



OM \rightarrow π bond ke 4-4 se foot do.

exothermic # ROR: Alkyne $>$ Alkene

4MCTs formed # ROR $\propto \frac{1}{\text{S.H}}$

Syn addition \propto stability



R_nNi \rightarrow alloy of Ni & Al in NaOH
Galvanic temp better than Ni (at 100°C)

If R_nNi used \rightarrow Sabatier's Sanderson's

If [Rh(PPh₃)₃Cl] used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	E ₂ g ₁ g ₂ g ₃
Trans	Syn	T ₂ e ₁
Cis	Anti	T ₂ e ₂
Trans	Anti	E ₂ g ₁ g ₂ g ₃

2. From alkyl halide



X hafao \propto leaque.

ROR: R-Cl $<$ R-Br $<$ R-I # ROR $\propto \frac{1}{\text{S.H}}$

3. Wurtz reaction



produce higher alkane with even no. of C

R-X \rightarrow coupling

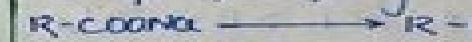
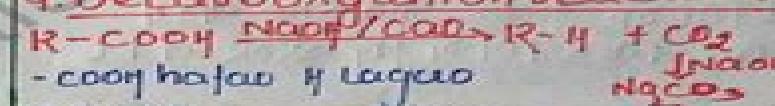
(1, 2nd)
R-X \rightarrow disproportionation

(3rd)
Intra molecular Wurtz faster than inter molecular but prod should be stable i.e. A-A nahi banana hai.



for same R: RF $<$ RCl $<$ RBr $<$ RI

4. Decarboxylation reaction



ROR \propto stability of intermediate (R⁻)

5. Kolbe's electrolysis



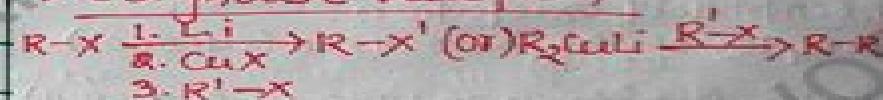
Cathode $\rightarrow \text{H}_2$ Anode \rightarrow Alkane + CO₂

Addition Reaction

6. Frankland reaction



7. Cosyhouse reaction



Attack of R⁻ \rightarrow RDS

ROR $\propto \frac{1}{\text{S.H}}$ of R'-X

If same R:



BE $\downarrow \rightarrow$ leaving tendency \uparrow

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Daniel C. Pantaleo

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1. Understanding the eBook Chemistry Study Guide Hydrocarbons Answers
 - The Rise of Digital Reading Chemistry Study Guide Hydrocarbons Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Hydrocarbons Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Hydrocarbons Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Hydrocarbons Answers
 - Personalized Recommendations
 - Chemistry Study Guide Hydrocarbons Answers User Reviews and Ratings
 - Chemistry Study Guide Hydrocarbons Answers and Bestseller Lists
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 - Chemistry Study Guide Hydrocarbons Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Chemistry Study Guide Hydrocarbons Answers
 - Interactive Elements Chemistry Study Guide Hydrocarbons Answers
- 8. Staying Engaged with Chemistry Study Guide Hydrocarbons Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Hydrocarbons Answers
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Hydrocarbons Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Hydrocarbons Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Hydrocarbons Answers
 - Setting Reading Goals Chemistry Study Guide Hydrocarbons Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Hydrocarbons Answers
 - Fact-Checking eBook Content of Chemistry Study Guide Hydrocarbons Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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