

**CHAPTER 17** | **STUDY GUIDE****Chemical Equilibrium****Section 17.1 Equilibrium: A State of Dynamic Balance***In your textbook, read about chemical equilibrium.***Complete each statement.**

1. When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to \_\_\_\_\_.
2. A reaction that can occur in both the forward and the reverse directions is called a(n) \_\_\_\_\_.
3. \_\_\_\_\_ is a state in which the forward and reverse reactions balance each other because they take place at equal rates.
4. At equilibrium, the concentrations of reactants and products are \_\_\_\_\_ but that does not mean that the amounts or concentrations are \_\_\_\_\_.
5. Equilibrium is a state of \_\_\_\_\_, not one of \_\_\_\_\_.

*In your textbook, read about equilibrium expressions and constants.***For each statement below, write true or false.**

\_\_\_\_\_ 6. The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value.

\_\_\_\_\_ 7. The equation  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \leftrightarrow 2\text{HI}(\text{g})$  is an example of a homogeneous equilibrium.

\_\_\_\_\_ 8. If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.

\_\_\_\_\_ 9. The value for  $K_{\text{eq}}$  is constant only at a specific volume.

\_\_\_\_\_ 10. If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.

\_\_\_\_\_ 11. A heterogeneous equilibrium means that reactants and products are present in more than one state.

\_\_\_\_\_ 12. The product of the forward chemical reaction is  $\text{HI}$ , for the equilibrium expression:

$$K_{\text{eq}} = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

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**Katherine J. Denniston**

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challenging problems in which the student must apply the skill in a slightly different environment. These problems include conceptual cumulative and applied problems that encourage students to think out of the box. Sometimes problems that foreshadow concepts introduced in later chapters are also included. All SkillBuilders are visually summarized at the end of each chapter. Skillbuilder review followed by a list of suggested in chapter and end of chapter practice problems. David Klein is a lecturer at Johns Hopkins University where he teaches Organic and General Chemistry. He is a dynamic and creative teacher and uses analogy to help students grasp difficult topics. Klein's unique informal voice and manner of presentation help students truly master key topics in this course. He is also the author of *Organic Chemistry as a Second Language* and *General Chemistry as a Second Language* which have both been highly successful.

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