

Lab Day 5, Recrystallization & Melting Points Prelab Questions**Prelab Questions**

1) During a melting point determination, the heating is constant and causes the temperature of the unit to continuously rise. In light of that, explain in your words why a measured melting point can never truly be a single temperature. Which would give a smaller melting point range, slower heating or more rapid heating? A small sample or a large sample? Explain.

Melting point can never truly be a single temperature because a solid melts over a range of temperatures instead of a specific temperature. Using a range would be better because it gives when the solid starts to melt and then when it finishes melting. Slower heating would give a smaller melting point range because it gives it a more precise range. A small sample would require less heat than a larger sample to melt therefore it would also give a smaller melting point range.

2) Explain why it is common practice to determine a melting point on an unknown compound rapidly first, and then to rerun the melting point slowly.

It is common practice to determine the melting point rapidly first so that you can determine the range so that you can get an estimate of where the boiling point is. The second time which is to rerun the melting point slowly will give you a more accurate reading.

3) This lab uses a mixed solvent system. Explain the possible advantages of using two solvents and the properties of those solvents that are important. Don't look up the answers, think.

The advantage of using two solvents is that it will help with the growth of the crystals.

4) If fine sand were an impurity in a sample, how would it effect the melting point? What would make you suspicious that something inert like sand was contaminating your sample?

If sand were an impurity in the sample, I think it would make the melting point high since it requires a lot of heat to melt sand. Something that would make me suspicious is if the sample is taking a while to melt even at a very high temperature.

Chem 121 Lab 5 Manual Answers

Elizabeth Kean, Catherine Middlecamp



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