

9. The IUPAC name of the compound  $[\text{Cu}(\text{NH}_3)_4](\text{NO}_3)_2$  is :  
(A) Cuprammonium nitrate (B) Tetraammine copper (II) dinitrate  
(C) Tetraammine copper (II) nitrate (D) Tetraammine copper (III) dinitrate
10. Hexafluoroferrate (III) ion is an outer orbital complex. The number of unpaired electrons present in it is :  
(A) 1 (B) 5 (C) 4 (D) Unpredictable.
11. Which of the following complex species involves  $d^2sp^3$  hybridisation :  
(A)  $[\text{CoF}_6]^{3-}$  (B)  $[\text{Co}(\text{H}_2\text{O})_6]^{3+}$  (C)  $[\text{Fe}(\text{CN})_6]^{3-}$  (D)  $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ .
12. How many ions are produced from  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2$  in solution :  
(A) 6 (B) 4 (C) 3 (D) 2.
13. In the complex  $\text{Fe}(\text{CO})_x$ , the value of x is :  
(A) 3 (B) 4 (C) 5 (D) 6.
14. Which is the central ion in  $[\text{Cu}(\text{H}_2\text{O})_4]^{2+}$  ion :  
(A)  $\text{Cu}^{2+}$  (B)  $\text{H}_3\text{O}^+$  (C)  $\text{Cu}^+$  (D) None.
15. In  $\text{K}_4[\text{Fe}(\text{CN})_6]$  the number of unpaired electrons in iron are :  
(A) 0 (B) 2 (C) 3 (D) 5.
16. All ligands are :  
(A) Lewis acid (B) Lewis base (C) Neutral (D) None.

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