

16 • Chemical Equilibrium

1. Consider the reaction system.
 $\text{CoO(s)} + \text{H}_2\text{g} \rightleftharpoons \text{Co(s)} + \text{H}_2\text{O(g)}$.

The equilibrium constant expression is

a) $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$ d) $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$
 b) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$ c) $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$
 (d) $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$
 (e) $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$

2. Given the equilibrium,
 $2\text{SO}_2\text{g} + \text{O}_2\text{g} \rightleftharpoons 2\text{SO}_3\text{g}$, if this equilibrium is established by beginning with equal number of moles of SO_2 and O_2 in a 1.0 Liter bulb, then the following *must* be true at equilibrium:
 a) $[\text{SO}_2] = [\text{SO}_3]$ d) $[\text{SO}_2] < [\text{O}_2]$
 (d) $2[\text{SO}_2] = 2[\text{SO}_3]$ e) $[\text{SO}_2] > [\text{O}_2]$
 c) $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO , 0.200 mol Cl_2 and 0.500 mol ClNO were placed in a 25.0 Liter container. The following equilibrium is established: $2\text{ClNO(g)} \rightleftharpoons 2\text{NO(g)} + \text{Cl}_2\text{g}$

3. At equilibrium, 0.600 mol of ClNO was present. The number of *moles* of Cl_2 present at equilibrium is
 (a) 0.050 (d) 0.200
 b) 0.100 e) 0.250
 c) 0.150

PRACTICE TEST

4. The equilibrium constant, K_c , is:
 a) 4.45×10^{-4} d) 0.167
 b) 6.67×10^{-4} e) 1500
 (c) 0.111

5. At 985°C , the equilibrium constant for the reaction,
 $\text{H}_2\text{g} + \text{CO}_2\text{g} \rightleftharpoons \text{H}_2\text{O(g)} + \text{CO(g)}$

is 1.63. What is the equilibrium constant for the reverse reaction?
 a) 1.63 d) 0.613
 b) 0.815 (c) 1.00
 c) 2.66

6. What is the relationship between K_p and K_c for the reaction, $2\text{ICl(g)} \rightleftharpoons \text{I}_2\text{g} + \text{Cl}_2\text{g}$?
 (a) $K_p = K_c(RT)^{-1}$ d) $K_p = K_c$
 b) $K_p = K_c(RT)$ e) $K_p = K_c(2RT)$
 c) $K_p = K_c(RT)^2$

7. For the reaction $2\text{NO}_2\text{g} \rightleftharpoons \text{N}_2\text{O}_4\text{g}$, K_p at 25°C is 7.3, when all partial pressures are expressed in atmospheres. What is K_c for this reaction? [$R=0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$]
 (a) 4270 d) 179
 b) 0.0119 e) 2.06
 c) 0.291

Chemical Equilibrium Test

**Donald Pritchard Smith, Henry Keppel
Miller**

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SAT Subject Test Chemistry Joseph A. Mascetta, Mark Kernion, 2020-12-01 Note College Board has discontinued the SAT Subject Tests in the US The tests will be available outside the US in June 2021 and

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Hazardous and Industrial Solid Waste Testing and Disposal Douglas Lorenzen, American Society for Testing and Materials, 1986

An Introduction to Qualitative Chemical Analysis and the Related Chemical Principles Donald Pritchard Smith, Henry Keppele Miller, 1928

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Bernard Ekeley, 1923

Tests of the Fatigue Strength of Cast Steel University of Illinois (Urbana-Champaign campus).

Engineering Experiment Station, Alfred Copeland Callen, Crandall Zachariah Rosecrans, Cloyd Moffett Smith, Herbert Fisher Moore, 1926

Laboratory Manual for General Chemistry Alvin Strickler, 1923

Laboratory Manual in General College

Chemistry Joseph Albert Babor, Alexander Lehrman, 1929

A Laboratory Manual in General Chemistry William Foster, 1923

Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2012-02-17

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Table of Contents Chemical Equilibrium Test

1. Understanding the eBook Chemical Equilibrium Test
 - The Rise of Digital Reading Chemical Equilibrium Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Equilibrium Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Test
 - Personalized Recommendations
 - Chemical Equilibrium Test User Reviews and Ratings
 - Chemical Equilibrium Test and Bestseller Lists
5. Accessing Chemical Equilibrium Test Free and Paid eBooks

- Chemical Equilibrium Test Public Domain eBooks
- Chemical Equilibrium Test eBook Subscription Services
- Chemical Equilibrium Test Budget-Friendly Options

6. Navigating Chemical Equilibrium Test eBook Formats

- ePUB, PDF, MOBI, and More
- Chemical Equilibrium Test Compatibility with Devices
- Chemical Equilibrium Test Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemical Equilibrium Test
- Highlighting and Note-Taking Chemical Equilibrium Test
- Interactive Elements Chemical Equilibrium Test

8. Staying Engaged with Chemical Equilibrium Test

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemical Equilibrium Test

9. Balancing eBooks and Physical Books Chemical Equilibrium Test

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemical Equilibrium Test

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemical Equilibrium Test

- Setting Reading Goals Chemical Equilibrium Test
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemical Equilibrium Test

- Fact-Checking eBook Content of Chemical Equilibrium Test
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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