

General Chemistry II Chapter 15 Study Guide

- $S'_{\text{gas}} > S'_{\text{liquid}} > S'_{\text{solid}}$
- For similar molecules, S' tends to increase with:
 - Increasing molar mass
 - Increased number of atoms in the formula

Standard entropy of a reaction ($\Delta S'_{\text{rxn}}$):

- $\Delta S'_{\text{rxn}}$: is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S'_{\text{rxn}} = \Sigma m S'_{\text{products}} - \Sigma n S'_{\text{reactants}}$$

where m and n are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction: $a A + b B \rightarrow c C + d D$

$$\Sigma m S'_{\text{products}} = c \Delta S'_{\text{rxn}} (C) + d \Delta S'_{\text{rxn}} (D)$$

$$\Sigma n S'_{\text{reactants}} = a \Delta S'_{\text{rxn}} (A) + b \Delta S'_{\text{rxn}} (B)$$

Remember to multiply ΔH° of each substance by its coefficient.

Predicting the sign of $\Delta S'_{\text{rxn}}$:

- We can often predict the sign of $\Delta S'_{\text{rxn}}$ for processes that involve a change in the number of moles of gas
 - $\Delta S'_{\text{rxn}}$ is positive if the amount of gas increases
 - $\Delta S'_{\text{rxn}}$ is negative if the amount of gas decreases
 - $\Delta S'_{\text{rxn}}$ is likely to be positive if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy (G): combines the enthalpy and entropy of a system:
 - $G = H - TS$
- Gibbs free energy change (ΔG): is a measure of the spontaneity of a process and of the useful energy available from it:
 - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$ for a spontaneous reaction ($-\Delta G$)
- $\Delta G > 0$ for a nonspontaneous reaction ($+\Delta G$)
- $\Delta G = 0$ for a process at equilibrium ($\Delta G = 0$)

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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end. Two full length practice tests with detailed answer explanations are included in the book.

Wastewater Operator Certification Study Guide John Giorgi, 2011-01-12 **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994 *Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject. The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors. Also of benefit is the just in time presentation of key content only providing details once they are needed. While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium. The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons. The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor's resource CD and a solution CD.

Study Guide to Chemistry and Life John William Hill, Dorothy M. Feigl, 1978 **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises. Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry. This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines *Chemistry, Study Guide* Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry. Worked examples now number over 300 and exercises over 1460 Introduction to Chemistry, Study Guide T. R. Dickson, 1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining problem solving skills. Many chapters in this edition have been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference. The Overview section summarizes the content of the chapter and

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