

## General Chemistry II Chapter 15 Study Guide

- $S_{\text{gas}} > S_{\text{liquid}} > S_{\text{solid}}$
- For similar molecules,  $S^\circ$  tends to increase with:
  - Increasing molar mass
  - Increased number of atoms in the formula

Standard entropy of a reaction ( $\Delta S_{\text{rxn}}^\circ$ ):

- $\Delta S_{\text{rxn}}^\circ$ : is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S_{\text{rxn}}^\circ = \sum m S_{\text{products}}^\circ - \sum n S_{\text{reactants}}^\circ$$

where  $m$  and  $n$  are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction:  $a A + b B \rightarrow c C + d D$

$$\sum m S_{\text{products}}^\circ = c \Delta S^\circ (C) + d \Delta S^\circ (D)$$

$$\sum n S_{\text{reactants}}^\circ = a \Delta S^\circ (A) + b \Delta S^\circ (B)$$

Remember to multiply  $\Delta H^\circ$  of each substance by its coefficient.

Predicting the sign of  $\Delta S_{\text{rxn}}^\circ$ :

- We can often predict the sign of  $\Delta S_{\text{rxn}}^\circ$  for processes that involve a change in the number of moles of gas
  - $\Delta S_{\text{rxn}}^\circ$  is positive if the amount of gas increases
  - $\Delta S_{\text{rxn}}^\circ$  is negative if the amount of gas decreases
  - $\Delta S_{\text{rxn}}^\circ$  is likely to be positive if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy ( $G$ ): combines the enthalpy and entropy of a system:
  - $G = H - TS$
- Gibbs free energy change ( $\Delta G$ ): is a measure of the spontaneity of a process and of the useful energy available from it:
  - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$  for a spontaneous reaction ( $-\Delta G$ )
- $\Delta G > 0$  for a nonspontaneous reaction ( $+\Delta G$ )
- $\Delta G = 0$  for a process at equilibrium ( $\Delta G = 0$ )

# Chemistry Chapter 15 Study Guide

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*CliffsNotes AP Chemistry* Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

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*General Chemistry Key Concise Review Guide Sixth Edition* Darrell D. Ebbing, 1999      **Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition** G. Marc Loudon, Joseph G. Stowell, 1995      Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988      Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed Robert K. Wismer, 1982

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**Organic Chemistry Study Guide & Problems Book** Brent L. Iverson, 1998 Companion to Organic chemistry 2nd ed by William H Brown Provides a detailed section by section overview of the major points covered in the text All of the problems from the text are reprinted here with detailed stepwise solutions

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