



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the reaction:

<p>What is Equilibrium?</p> <p>Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the reaction:</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is defined as:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant</p> <p>The equilibrium constant is a measure of the extent to which a chemical reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is defined as:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	 <p>Graphical Representation of Equilibrium:</p> <p>The graph shows concentration on the y-axis and time on the x-axis. Two curves are plotted: one for reactants and one for products. The reactant curve starts at a high concentration and decreases over time, while the product curve starts at zero and increases over time. Both curves eventually level off at the same time, indicating that the system has reached equilibrium. At equilibrium, the concentration of reactants is higher than the concentration of products.</p> <p>Reactants: $[A]_0, [B]_0, [C]_0, [D]_0$</p> <p>Products: $[A]_{eq}, [B]_{eq}, [C]_{eq}, [D]_{eq}$</p>												
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant is a measure of the extent to which a chemical reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is defined as:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_c:</p> <table><tr><th>Equilibrium Constant, K_c</th><th>Meaning</th></tr><tr><td>$K_c > 1$</td><td>Concentration of products is greater than concentration of reactants.</td></tr><tr><td>$K_c = 1$</td><td>Concentration of products is equal to concentration of reactants.</td></tr><tr><td>$K_c < 1$</td><td>Concentration of reactants is greater than concentration of products.</td></tr></table>		Equilibrium Constant, K_c	Meaning	$K_c > 1$	Concentration of products is greater than concentration of reactants.	$K_c = 1$	Concentration of products is equal to concentration of reactants.	$K_c < 1$	Concentration of reactants is greater than concentration of products.				
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<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p> <table><tr><th>Condition</th><th>Effect on Equilibrium</th><th>Example</th></tr><tr><td>Concentration</td><td><ul style="list-style-type: none">Increasing concentration of a reactant will shift the equilibrium to the right.Increasing concentration of a product will shift the equilibrium to the left.</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$<p>Increasing concentration of N_2 will shift the equilibrium to the right, producing more NH_3.</p></td></tr><tr><td>Pressure</td><td><ul style="list-style-type: none">Increasing pressure will shift the equilibrium to the side with fewer moles of gas.Decreasing pressure will shift the equilibrium to the side with more moles of gas.</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$<p>Increasing pressure will shift the equilibrium to the right, producing more NH_3.</p></td></tr><tr><td>Temperature</td><td><ul style="list-style-type: none">Increasing temperature will shift the equilibrium to the side that absorbs heat (endothermic reaction).Decreasing temperature will shift the equilibrium to the side that releases heat (exothermic reaction).</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$<p>Increasing temperature will shift the equilibrium to the left, producing less NH_3.</p></td></tr></table>			Condition	Effect on Equilibrium	Example	Concentration	<ul style="list-style-type: none">Increasing concentration of a reactant will shift the equilibrium to the right.Increasing concentration of a product will shift the equilibrium to the left.	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing concentration of N_2 will shift the equilibrium to the right, producing more NH_3.</p>	Pressure	<ul style="list-style-type: none">Increasing pressure will shift the equilibrium to the side with fewer moles of gas.Decreasing pressure will shift the equilibrium to the side with more moles of gas.	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing pressure will shift the equilibrium to the right, producing more NH_3.</p>	Temperature	<ul style="list-style-type: none">Increasing temperature will shift the equilibrium to the side that absorbs heat (endothermic reaction).Decreasing temperature will shift the equilibrium to the side that releases heat (exothermic reaction).	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing temperature will shift the equilibrium to the left, producing less NH_3.</p>
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Chemistry Equilibrium Study Guide

Thomas Elliott Taylor



Chemistry Equilibrium Study Guide:

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Fundamentals of Chemistry, Study Guide* James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs *Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth

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Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed Robert K. Wismer, 1982

General Chemistry, Study Guide James E. Brady, 1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

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Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

Resources in Education, 1984

Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

Study Guide for Chemical Principles Thomas Elliott Taylor, 1979

This book delves into Chemistry Equilibrium Study Guide. Chemistry Equilibrium Study Guide is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemistry Equilibrium Study Guide, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 2: Essential Elements of Chemistry Equilibrium Study Guide
 - Chapter 3: Chemistry Equilibrium Study Guide in Everyday Life
 - Chapter 4: Chemistry Equilibrium Study Guide in Specific Contexts
 - Chapter 5: Conclusion
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