

ATOMIC STRUCTURE

RUTHERFORD'S MODEL

- Discovered that an atom is mostly empty space with tiny particles.
- Charged particles move in a small elliptical path around a central nucleus.
- This can only hold the negatively charged particles in the center.

BOHR'S MODEL

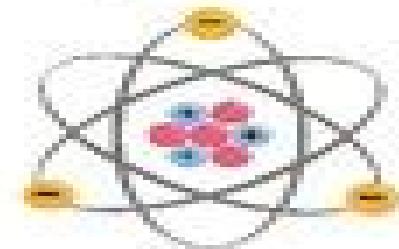
- Shows the elliptical path of an electron.
- Electrons
- Electrons

THE JELLY PUDDING MODEL



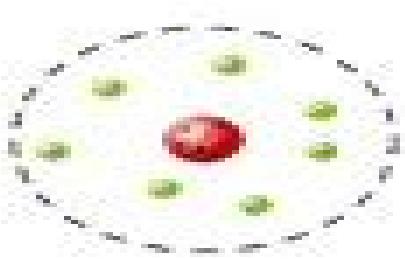
WAVE-MECHANICAL MODEL

- Shows electrons moving in paths, like tiny clouds like in clouds. Orbitals are regions of space likely where an electron.



WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
 - PROTONS - positively charged particles
 - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
 - Electrons DO move around the nucleus in different patterns.



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1.67 x 10 ⁻²⁴
neutron	center (in the nucleus)	neutral (no charge)	1.67 x 10 ⁻²⁴
electron	around the nucleus in elliptical orbits	negative	1.67 x 10 ⁻²⁴ x 10 ⁻³⁰

IONS

- If an atom gains or loses an electron, it becomes an ion.
- This gives the atom a charge:
 - Loss of electrons, the atom becomes positive.
 - Gains of electrons, the atom becomes negative.

ISOTOPES

- An isotope is when an atom has a different mass number.
- An isotope is still the same element, it is just a little different mass, every atom has all other atoms different in slightly different mass.

Chemistry Atomic Structure Study Guide

Samuel C. Brownstein, Mitchel Weiner

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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy,2011 Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons,Craig B. Fryhle,Scott A. Snyder,2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993

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