

10**CHEMICAL QUANTITIES****Practice Problems***On a piece of paper, solve the following problems.****YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!*****SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

- What is the molar mass of sucrose ($C_{12}H_{22}O_{11}$)?
- What is the molar mass of each of the following compounds?
 - phosphorus pentachloride (PCl_5)
 - uranium hexafluoride (UF_6)
- Calculate the molar mass of each of the following ionic compounds:
 - $KMnO_4$
 - $Ca_3(PO_4)_2$
- How many moles is 3.52×10^{24} molecules of water?
- How many atoms of zinc are in 0.60 mol of zinc?
- What is the mass of 1.00 mol of oxygen (O_2)?

SECTION 10.2 MOLE-MASS AND MOLE-VOLUME RELATIONSHIPS

- What is the molar mass of each of the following compounds?
 - $C_6H_{12}O_6$
 - $NaHCO_3$
 - C_2H_6
 - KNH_4SO_4
- Calculate the mass in grams of each of the following:
 - 8.0 mol lead oxide (Pb_3O_4)
 - 1.50 $\times 10^{-2}$ mol molecular oxygen (O_2)
 - 0.75 mol hydrogen sulfide (H_2S)
 - 2.30 mol ethylene glycol ($C_2H_6O_2$)
 - 0.00100 mol silicon tetrahydride (SiH_4)
- How many grams are in 1.73 mol of dinitrogen pentoxide (N_2O_5)?
- How many grams are in 0.658 mol of calcium phosphate [$Ca_3(PO_4)_2$]?
- Calculate the number of moles in each of the following:
 - 0.50 g sodium bromide ($NaBr$)
 - 0.00100 g monochloromethane (CH_3Cl)
 - 13.5 g magnesium nitrate [$Mg_2(NO_3)_3$]
 - 1.50 $\times 10^{-3}$ g propylene glycol [$C_3H_6(OH)_2$]
 - 1.02 g magnesium chloride ($MgCl_2$)
- A chemist plans to use 435.0 grams of ammonium nitrate (NH_4NO_3) in a reaction. How many moles of the compound is this?
- A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ($C_{20}H_{24}N_4O_2$). What mass, in grams, should the laboratory technician obtain in order to make the solution?
- What is the volume at STP of 2.66 mol of methane (CH_4) gas?
- How many moles is 13.5 L of ammonia (NH_3) gas at STP?

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Answers

**Robert C. Smoot,Smoot,Richard G.
Smith,Jack Price**

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