

Chapter 13

Hydrocarbons

Question 13.1:

How do you account for the formation of ethane during chlorination of methane?

Answer 13.1:

Chlorination of methane proceeds via a free radical chain mechanism. The whole reaction takes place in the given three steps.

Step 1: Initiation:

The reaction begins with the homolytic cleavage of Cl – Cl bond as:



Step 2: Propagation:

In the second step, chlorine free radicals attack methane molecules and break down the C–H bond to generate methyl radicals as:



These methyl radicals react with other chlorine free radicals to form methyl chloride along with the liberation of a chlorine free radical.



Hence, methyl free radicals and chlorine free radicals set up a chain reaction. While HCl and CH_3Cl are the major products formed, other higher halogenated compounds are also formed as:



Step 3: Termination:

Formation of ethane is a result of the termination of chain reactions taking place as a result of the consumption of reactants as:



Hence, by this process, ethane is obtained as a by-product of chlorination of methane.

Chemistry Chapter 13 Solutions Manual

Carl Schorlemmer

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Instructor's manual and solutions to problems to accompany General chemistry: principles and structures James E. Brady, 1975-04 Solutions Manual to Accompany Chemistry Raymond Chang, 1984 **Student Solutions Manual to Accompany Chemistry** William Joseph Pietro, 1990 **Student Solutions Manual to Accompany Chemistry & Chemical Reactivity, Fourth Edition, Kotz & Treichel** Alton J. Banks, Paul Treichel, 1999 Atkins' Physical Chemistry Peter Atkins, Julio de Paula, 2010 This volume features a greater emphasis on the molecular view of physical chemistry and a move away from classical thermodynamics. It offers greater explanation and support in mathematics which remains an intrinsic part of physical chemistry.

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