

Chapter 13

Hydrocarbons

Question 13.1:

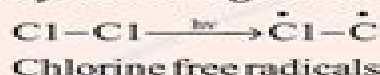
How do you account for the formation of ethane during chlorination of methane?

Answer 13.1:

Chlorination of methane proceeds via a free radical chain mechanism. The whole reaction takes place in the given three steps.

Step 1: Initiation:

The reaction begins with the homolytic cleavage of Cl – Cl bond as:



Step 2: Propagation:

In the second step, chlorine free radicals attack methane molecules and break down the C–H bond to generate methyl radicals as:



These methyl radicals react with other chlorine free radicals to form methyl chloride along with the liberation of a chlorine free radical.



Hence, methyl free radicals and chlorine free radicals set up a chain reaction. While HCl and CH₃Cl are the major products formed, other higher halogenated compounds are also formed as:



Step 3: Termination:

Formation of ethane is a result of the termination of chain reactions taking place as a result of the consumption of reactants as:



Hence, by this process, ethane is obtained as a by-product of chlorination of methane.

Chemistry Chapter 13 Solutions Manual

Carl Schorlemmer



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the important material of the course and explain it in a concise clear way
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A Manual of the Chemistry of the Carbon

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Table of Contents Chemistry Chapter 13 Solutions Manual

1. Understanding the eBook Chemistry Chapter 13 Solutions Manual
 - The Rise of Digital Reading Chemistry Chapter 13 Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 13 Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 13 Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 13 Solutions Manual
 - Personalized Recommendations

- Chemistry Chapter 13 Solutions Manual User Reviews and Ratings
- Chemistry Chapter 13 Solutions Manual and Bestseller Lists
- 5. Accessing Chemistry Chapter 13 Solutions Manual Free and Paid eBooks
 - Chemistry Chapter 13 Solutions Manual Public Domain eBooks
 - Chemistry Chapter 13 Solutions Manual eBook Subscription Services
 - Chemistry Chapter 13 Solutions Manual Budget-Friendly Options
- 6. Navigating Chemistry Chapter 13 Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 13 Solutions Manual Compatibility with Devices
 - Chemistry Chapter 13 Solutions Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 13 Solutions Manual
 - Highlighting and Note-Taking Chemistry Chapter 13 Solutions Manual
 - Interactive Elements Chemistry Chapter 13 Solutions Manual
- 8. Staying Engaged with Chemistry Chapter 13 Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 13 Solutions Manual
- 9. Balancing eBooks and Physical Books Chemistry Chapter 13 Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 13 Solutions Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 13 Solutions Manual
 - Setting Reading Goals Chemistry Chapter 13 Solutions Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 13 Solutions Manual
 - Fact-Checking eBook Content of Chemistry Chapter 13 Solutions Manual

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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