

CHAPTER 14 STUDY GUIDE

Mixtures and Solutions

Section 14.1 Heterogeneous and Homogeneous Mixtures

In your textbook, read about suspensions and colloids.

For each statement below, write *true* or *false*.

- | | |
|-------|---|
| false | 1. A solution is a mixture containing particles that settle out of the mixture if left undisturbed. |
| true | 2. The most abundant substance in a colloid is the dispersion medium. |
| false | 3. A colloid can be separated by filtration. |
| true | 4. A solid emulsion consists of a liquid dispersed in a solid. |
| true | 5. Whipped cream is an example of a foam. |
| false | 6. In an aerosol, the dispersing medium is a liquid. |
| true | 7. Brownian motion results from the collisions of particles of the dispersion medium with the dispersed particles. |
| true | 8. Dispersed particles in a colloid do not tend to settle out because they have polar or charged atomic groups on their surfaces. |
| false | 9. Stirring an electrolyte into a colloid stabilizes the colloid. |
| true | 10. Colloids demonstrate the Tyndall effect. |

The table below lists the characteristics of particles in colloids, solutions, and suspensions. Place a check in the column of each mixture whose particles have a particular characteristic.

Characteristics of Particles	Colloid	Solution	Suspension
11. Less than 1 nm in diameter		✓	✓
12. Between 1 nm and 1000 nm in diameter	✓	✓	✓
13. More than 1000 nm in diameter	✓	✓	
14. Settle out if undisturbed			
15. Pass through standard filter paper	✓		
16. Lower vapor pressure			✓
17. Scatter light	✓		

Chemistry Chapter 14 Mixtures Solutions Study Guide Answers

Darrell D. Ebbing, Susan M. Schelble



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Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

Study Guide/Selected Solutions Manual Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

CHEMISTRY SILBERBERG, 2003 Know Your 'O' Level Chemistry - A Study Guide, Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006

Study Guide Joan I. Senyk, Darrell D. Ebbing, 1984 **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W.

Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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Ebbing, Susan M. Schelble, 1998 Study Guide to Chemistry and Life John William Hill, Dorothy M. Feigl, 1978 Study Guide [to Accompany] General Chemistry Dixie Goss, 2002 Im/Tb Chemistry F/Today Slabaugh, Seager, 2004 **Biology**
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