

Chapter 15 - **KEY** Solutions

Solution - homogeneous mixture containing two or more substances called the solute and the solvent

A solution may exist as a gas (air), liquid (sugar and water), or solid (braces – titanium and nickel).
Most solutions are liquid.

Aqueous solutions - reactants and products that are mixed in water.

Water is the most common solvent among liquid solutions. It is known as the "universal solvent."

soluble - a substance that dissolves in a substance.

Ex. -sugar is soluble in water.

insoluble - a substance that does not dissolve in a solvent.

Ex. - sand is insoluble in water.

immiscible - 2 liquids that do not dissolve each other Ex. - oil and vinegar

miscible - 2 liquids that are soluble in each other Ex. - food coloring and water

solvation - the process of surrounding solute particles with solvent particles to form a solution.

If the attractive forces between a solute and a solvent are greater than the attractive forces holding the solute together, then the solvent particles pull the solute particles apart and surround them.

General rule is that "like dissolves like" (polar dissolves polar, nonpolar dissolves nonpolar, etc.)

Ex. – water and NaCl

Factors that affect rate of solvation –

1. agitation - moves the solute particles around and thereby allows for more collisions
2. Increase the surface area of the solute -this also allows more collisions to occur (tablespoon of sugar dissolves faster than a cube).
3. raising the temperature of the solvent -this increases the kinetic energy of its particles, resulting in more frequent collisions.

Heat of Solution - the overall energy change that occurs during the solvation process.

Energy is required to overcome the attractive forces within the solute and solvent (endothermic)

When solute and solvent mix, the particles attract each other and release energy (exothermic)

Solubility - maximum amount of solute that will dissolve in a given amount of solvent at a specified temp. and pressure.

Saturated solution - contains the maximum amount of dissolved solute for a given amount of solvent at a specific temp. and pressure

Unsaturated Solution - contains less dissolved solute for a given temp. and pressure than a saturated solution

Chemistry Chapter 15 Solutions Manual

John Williams

Chemistry Chapter 15 Solutions Manual:

Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter,Charles Trapp,Peter Atkins,Loretta Jones,1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools **Solutions Manual to Accompany Chemistry** Raymond Chang,1984 Solutions Manual with Answers to All Questions, Analytical Chemistry, Principles and Techniques Larry G. Hargis,1987-12 Solutions Manual to Accompany Physical Chemistry Ira N. Levine,1995 **Student Solutions Manual, Chemistry, Principles & Practice, Second Edition** John M. DeKorte,1997 *Student Guide and Solutions Manual to Accompany Ternay's Contemporary Organic Chemistry* Robert F. Francis,1979 *Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines **Student Solutions Manual for Chang's Chemistry** Philip C. Keller,Raymond Chang,Jill L. Keller,1988

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