

Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

Chemical Bonding:

- **Four types of bonds:**
 - *Nonpolar Covalent Bond* (ΔEN (difference in electronegativity) = < 0.3)
 - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
 - *Polar Covalent Bond* ($\Delta EN = 0.4-1.7$)
 - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
 - *Ionic Bond* ($\Delta EN = > 1.7$)
 - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
 - Ionic bond results due to the attraction between the **POSITIVE** cation (formed by losing electrons) and the **NEGATIVE** anion (formed by gaining electrons)
 - *Metallic Bond*
 - Bonding between two metallic atoms which occurs due to the attraction between a "sea of electrons" and the nuclei of bonded metal atoms

* Must memorize/know the difference in electronegativity ranges for each type of bond

** Note: you should still know the definition of electronegativity (EN)

- **Properties of the Four Types of Chemical Bonds**
 - See the table at the bottom of Packet pg.5
 - You **MUST** be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
 - Why → because all atoms want to be at their **LOWEST POTENTIAL ENERGY**; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
 - How → **ATTRACTIVE** and **REPULSIVE** forces
 - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
 - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repeal one another (because they have the same charge)
 - *Bond Length*
 - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
 - *Bond Energy*
 - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

Chemistry Chapter 6 Study Guide

Karen Timberlakes

Chemistry Chapter 6 Study Guide:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Hazmat Chemistry Study Guide (Second Edition)* Jill Meryl Levy,2011

A Self-study Guide to the Principles of Organic Chemistry Jiben Roy,2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix

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