

Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

Chemical Bonding:

- **Four types of bonds:**

- *Nonpolar Covalent Bond* (ΔEN (difference in electronegativity) = < 0.3)
 - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
- *Polar Covalent Bond* ($\Delta EN = 0.4-1.7$)
 - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
- *Ionic Bond* ($\Delta EN = > 1.7$)
 - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
 - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
- *Metallic Bond*
 - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms

* Must memorize/know the difference in electronegativity ranges for each type of bond

** Note: you should still know the definition of electronegativity (EN)

- **Properties of the Four Types of Chemical Bonds**

- See the table at the bottom of Packet pg.5
- You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness

- **Why & How do chemical bonds form?**

- **Why** → because all atoms want to be at their LOWEST POTENTIAL ENERGY; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
- **How** → ATTRACTIVE and REPULSIVE forces
 - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
 - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repeal one another (because they have the same charge)
- **Bond Length**
 - Ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
- **Bond Energy**
 - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

Chemistry Chapter 6 Study Guide

Karen Timberlakes



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Hazmat Chemistry Study Guide (Second Edition)* Jill Meryl Levy, 2011 **A Self-study Guide to the Principles of Organic Chemistry** Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix **Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed** Robert K. Wismer, 1982 **Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988 **Study**

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Table of Contents Chemistry Chapter 6 Study Guide

1. Understanding the eBook Chemistry Chapter 6 Study Guide
 - The Rise of Digital Reading Chemistry Chapter 6 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 6 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 6 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 6 Study Guide
 - Personalized Recommendations

- Chemistry Chapter 6 Study Guide User Reviews and Ratings
- Chemistry Chapter 6 Study Guide and Bestseller Lists
- 5. Accessing Chemistry Chapter 6 Study Guide Free and Paid eBooks
 - Chemistry Chapter 6 Study Guide Public Domain eBooks
 - Chemistry Chapter 6 Study Guide eBook Subscription Services
 - Chemistry Chapter 6 Study Guide Budget-Friendly Options
- 6. Navigating Chemistry Chapter 6 Study Guide eBook Formats
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 - Chemistry Chapter 6 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 6 Study Guide
 - Highlighting and Note-Taking Chemistry Chapter 6 Study Guide
 - Interactive Elements Chemistry Chapter 6 Study Guide
- 8. Staying Engaged with Chemistry Chapter 6 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 6 Study Guide
- 9. Balancing eBooks and Physical Books Chemistry Chapter 6 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 6 Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 6 Study Guide
 - Setting Reading Goals Chemistry Chapter 6 Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 6 Study Guide
 - Fact-Checking eBook Content of Chemistry Chapter 6 Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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