

Unit Three Study Guide -- Chemistry

Answer the following questions to prepare for your next test. The completed study guide is due Test Day (Friday October 3). Be sure to get extra help if you struggle on any of the concepts in this study guide. All answers should be on a separate sheet of paper.

1. Define the following words.

- a. atomic size
- b. ionization energy
- c. ions
- d. cations
- e. anions

- f. ionic compounds
- g. covalent compounds
- h. electron configuration
- i. valence electrons
- j. energy levels

2. Write the FULL electron configuration for the following elements.

- a. Sodium
- b. Potassium
- c. Rubidium
- d. Vanadium
- e. Helium
- f. Cobalt

- g. Gallium
- h. Argon
- i. Chlorine
- j. Carbon
- k. Bromine
- l. Krypton

3. Identify the element represented by the following electron configurations.

- a. $1s^2 2s^1$
- b. $1s^2 2s^2 2p^6 3s^1$
- c. $1s^2 2s^2 2p^2$
- d. $1s^2 2s^2 2p^6 3s^2 3p^1$
- e. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^1$
- f. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$
- g. $1s^2 2s^2 2p^2$

- h. $[Ar] 4s^2 3d^{10}$
- i. $[He] 2s^2 2p^2$
- j. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^2$
- k. $[Xe] 6s^2 5d^1$
- l. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^2$
- m. $[Kr] 5s^1$
- n. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^1$

4. Identify whether the trend is increasing, decreasing, or stays the same.

- a. electronegativity as you move left to right across the table
- b. ionization energy as you move down the table
- c. electronegativity as you move down the table
- d. atomic radius as you move left to right on the table
- e. atomic radius as you move down the table
- f. ionization energy as you move left to right across the table
- g. atomic size as you move down the periodic table
- h. atomic size as you move across (left to right) the periodic table

5. Complete the table below.

Compound Formula	Ionic or Covalent?	Compound Name
H_2O		

Chemistry Chemical Bonding Study Guide Answers

Darrell D. Ebbing



Chemistry Chemical Bonding Study Guide Answers:

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Organic Chemistry Marc Loudon, Jim Parise, 2015-07-01 Loudon and Parise s Organic Chemistry is known for its clear writing high standard of accuracy and creative problems This edition contains over 1 800 problems many of them new and taken directly from the scientific literature The book is used at a wide variety of schools such as UC Berkeley Caltech Colorado Cornell Duke Harvard Illinois Maryland Purdue Yale Wisconsin and many more This edition provides students with more health examples drawn from modern medical practice as well as many cutting edge topics from modern synthetic organic chemistry In addition to the printed book students can rely on Sapling Learning s online homework platform for extra learning and assessment The platform offers automatic grading an easy to use interface and instructive feedback Instructors can select from a variety of existing problem sets over 1 000 of Loudon s problems are in the platform or they can modify the questions or author them from scratch Not only does the software allow students to easily draw and interact with structures it allows them to draw entire reaction mechanisms including showing the movement of electrons with curved electron arrows

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