

Unit Three: Study Guide --- Chemistry

Answer the following questions to prepare for your next test. The completed study guide is due Test Day (Friday October 30). Be sure to get extra help if you struggle on any of the concepts in this study guide. All answers should be on a separate sheet of paper.

1. Define the following words.

- a. atomic size
- b. ionization energy
- c. ions
- d. cations
- e. anions
- f. ionic compounds
- g. covalent compounds
- h. electron configurations
- i. valence electrons
- j. energy levels

2. Write the FULL electron configurations for the following elements.

- a. Sodium
- b. Potassium
- c. Rubidium
- d. Neon
- e. Helium
- f. Cobalt
- g. Gallium
- h. Argon
- i. Chlorine
- j. Carbon
- k. Boron
- l. Nitrogen

3. Identify the elements represented by the following electron configurations.

- a. $1s^2 2s^1$
- b. $1s^2 2s^1 2p^1 3s^1$
- c. $1s^2 2s^2 2p^1$
- d. $1s^2 2s^1 2p^1 3s^1 3p^1$
- e. $1s^2 2s^1 2p^1 3s^1 3p^1 4s^1 3d^1$
- f. $1s^2 2s^1 2p^1 3s^1 3p^1 4s^1 3d^1 4p^1$
- g. $1s^2 2s^1 2p^1 3s^1 3p^1 4s^1 3d^1$
- h. $1s^2 2s^1 2p^1 3s^1 3p^1 4s^1 3d^1 4p^1$
- i. $1s^2 2s^1 2p^1 3s^1 3p^1 4s^1$
- j. $1s^2 2s^1 2p^1 3s^1 3p^1 4s^1 3d^1$

4. Identify whether the trend is increasing, decreasing, or stays the same.

- a. ionization energy as you move left to right across the table
- b. ionization energy as you move down the table
- c. ionization energy as you move down the table
- d. atomic radius as you move left to right on the table
- e. atomic radius as you move down the table
- f. ionization energy as you move left to right across the table
- g. atomic size as you move down the periodic table
- h. atomic size as you move across (left to right) the periodic table

5. Complete the table below.

Compound Formula	Ionic or Covalent?	Compound Name
NaCl		

Chemistry Chemical Bonding Study Guide Answers

Darrell D. Ebbing

Chemistry Chemical Bonding Study Guide Answers:

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[by] Steven S. Zumdahl Paul B. Kelter, 1993

General Chemistry Study Guide Sixth Edition Darrell D. Ebbing, 1999

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