



# CHEMISTRY

## CHEAT SHEET

# EQUILIBRIUM

## EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol  $\rightleftharpoons$ .

<p><b>What is Equilibrium?</b></p> <p>Equilibrium is a state of balance between two or more opposing processes. In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p><b>Equilibrium Constant</b></p> <p>The equilibrium constant, <math>K</math>, is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Graphical Representation of Equilibrium</b></p> <p>The graph shows the concentration of reactants and products over time. The reactant concentration starts high and decreases exponentially towards a constant value. The product concentration starts at zero and increases exponentially towards a constant value. The system reaches equilibrium when the concentrations no longer change.</p>												
<p><b>Equilibrium Constant, <math>K_c</math></b></p> <p><math>K_c</math> is the equilibrium constant based on concentrations. It is used for reactions in solution or in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Equilibrium Constant, <math>K_p</math></b></p> <p><math>K_p</math> is the equilibrium constant based on partial pressures. It is used for reactions involving gases.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	<p><b>Equilibrium Constant, <math>K_c</math> vs. <math>K_p</math></b></p> <p>The relationship between <math>K_c</math> and <math>K_p</math> is given by the equation:</p> $K_p = K_c (RT)^{\Delta n}$ <p>where <math>R</math> is the gas constant, <math>T</math> is the temperature in Kelvin, and <math>\Delta n</math> is the change in the number of moles of gas.</p>												
<p><b>Equilibrium Constant, <math>K_a</math></b></p> <p><math>K_a</math> is the acid dissociation constant. It is used to measure the strength of an acid in solution.</p> <p>For a general acid dissociation reaction:</p> $HA \rightleftharpoons H^+ + A^-$ <p>The equilibrium constant is given by:</p> $K_a = \frac{[H^+] [A^-]}{[HA]}$	<p><b>Equilibrium Constant, <math>K_b</math></b></p> <p><math>K_b</math> is the base dissociation constant. It is used to measure the strength of a base in solution.</p> <p>For a general base dissociation reaction:</p> $B + H_2O \rightleftharpoons BH^+ + OH^-$ <p>The equilibrium constant is given by:</p> $K_b = \frac{[BH^+] [OH^-]}{[B] [H_2O]}$	<p><b>Equilibrium Constant, <math>K_a</math> vs. <math>K_b</math></b></p> <p>The relationship between <math>K_a</math> and <math>K_b</math> for a conjugate acid-base pair is given by the equation:</p> $K_a \cdot K_b = K_w$ <p>where <math>K_w</math> is the ion-product constant for water, which is <math>1.0 \times 10^{-14}</math> at 25°C.</p>												
<p><b>Le Chatelier's Principle: the effect of changing conditions on equilibrium</b></p> <table border="1"> <thead> <tr> <th>Condition</th> <th>Effect on Equilibrium</th> <th>Example</th> </tr> </thead> <tbody> <tr> <td>Concentration</td> <td> <ul style="list-style-type: none"> <li>Increasing the concentration of a reactant shifts the equilibrium to the right.</li> <li>Increasing the concentration of a product shifts the equilibrium to the left.</li> </ul> </td> <td> <math display="block">N_2 + 3H_2 \rightleftharpoons 2NH_3</math> <p>Increasing the concentration of <math>N_2</math> shifts the equilibrium to the right, producing more <math>NH_3</math>.</p> </td> </tr> <tr> <td>Pressure</td> <td> <ul style="list-style-type: none"> <li>Increasing the pressure shifts the equilibrium to the side with fewer moles of gas.</li> <li>Decreasing the pressure shifts the equilibrium to the side with more moles of gas.</li> </ul> </td> <td> <math display="block">N_2 + 3H_2 \rightleftharpoons 2NH_3</math> <p>Increasing the pressure shifts the equilibrium to the right, producing more <math>NH_3</math>.</p> </td> </tr> <tr> <td>Temperature</td> <td> <ul style="list-style-type: none"> <li>Increasing the temperature shifts the equilibrium to the endothermic side.</li> <li>Decreasing the temperature shifts the equilibrium to the exothermic side.</li> </ul> </td> <td> <math display="block">N_2 + 3H_2 \rightleftharpoons 2NH_3</math> <p>The forward reaction is exothermic. Increasing the temperature shifts the equilibrium to the left, producing less <math>NH_3</math>.</p> </td> </tr> </tbody> </table>			Condition	Effect on Equilibrium	Example	Concentration	<ul style="list-style-type: none"> <li>Increasing the concentration of a reactant shifts the equilibrium to the right.</li> <li>Increasing the concentration of a product shifts the equilibrium to the left.</li> </ul>	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of <math>N_2</math> shifts the equilibrium to the right, producing more <math>NH_3</math>.</p>	Pressure	<ul style="list-style-type: none"> <li>Increasing the pressure shifts the equilibrium to the side with fewer moles of gas.</li> <li>Decreasing the pressure shifts the equilibrium to the side with more moles of gas.</li> </ul>	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure shifts the equilibrium to the right, producing more <math>NH_3</math>.</p>	Temperature	<ul style="list-style-type: none"> <li>Increasing the temperature shifts the equilibrium to the endothermic side.</li> <li>Decreasing the temperature shifts the equilibrium to the exothermic side.</li> </ul>	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>The forward reaction is exothermic. Increasing the temperature shifts the equilibrium to the left, producing less <math>NH_3</math>.</p>
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# Chemistry Chemical Kinetics And Equilibrium Study Guide

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