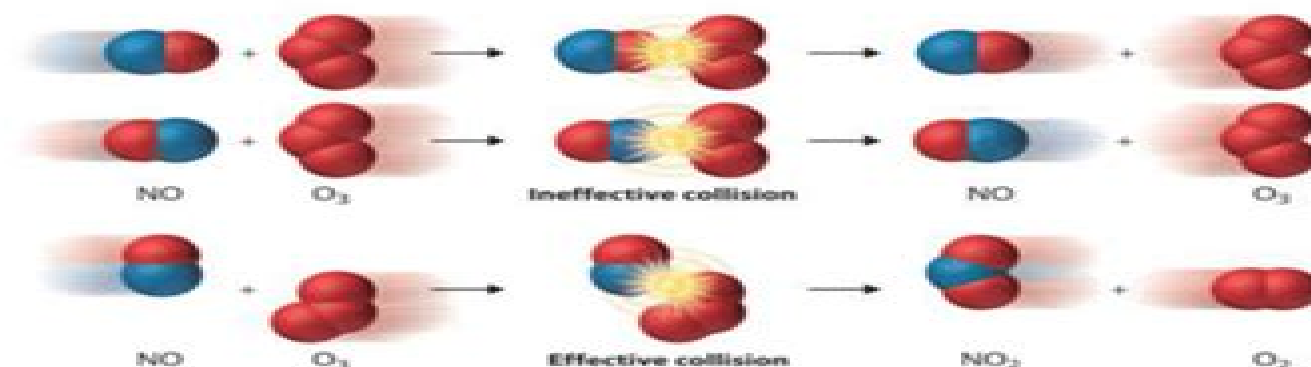


Topic	Collision theory	Level	GCSE
Outcomes	1. To use the collision theory to explain why changes to a reaction can increase the rate of reaction (temperature, concentration, surface area and catalyst) 2. To understand why not all collisions are successful		

### The collision theory and rates of reaction



*When two particles collide (hit each other) sometimes a chemical reaction happens. Not all collisions result in a chemical reaction. To increase the number of successful collisions we can change the conditions of a reaction (see page 2).*

1. Do all collisions between particles cause a reaction to happen?
2. What do we call the minimum energy the particles must have for a reaction to happen?
3. What do we call a collision between two particles that causes a chemical reaction?
4. Look at the picture above, how do you know the first reaction is not a successful collision?

# Chemistry Collision Theory Study Guide Answers

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