

Chemistry data sheet

Formula sheet

$$n = \frac{q}{zF}$$

$$q = mC\Delta T$$

$$pK_a = -\log_{10}[K_a]$$

$$c = \frac{Q}{V}$$

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$A = \epsilon lc = -\log_{10} \frac{I}{I_0}$$

$$PV = nRT$$

$$pH = -\log_{10}[H^+]$$

$$\text{Avogadro constant, } N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$$

$$\begin{aligned} \text{Volume of 1 mole ideal gas: at 100 kPa and} \\ \text{at } 0^\circ\text{C (273.15 K)} &= 22.71 \text{ L} \\ \text{at } 25^\circ\text{C (298.15 K)} &= 24.79 \text{ L} \end{aligned}$$

$$\text{Ionisation constant for water at } 25^\circ\text{C (298.15 K), } K_w = 1.0 \times 10^{-14}$$

$$\text{Specific heat capacity of water} = 4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$$

Solubility constants at 25°C

Compound	K_{sp}	Compound	K_{sp}
Barium carbonate	2.58×10^{-9}	Lead (II) bromide	6.60×10^{-9}
Barium hydroxide	2.55×10^{-4}	Lead (II) chloride	1.70×10^{-5}
Barium phosphate	1.3×10^{-29}	Lead (II) iodide	9.8×10^{-9}
Barium sulfate	1.08×10^{-10}	Lead (II) carbonate	7.4×10^{-14}
Calcium carbonate	3.36×10^{-9}	Lead (II) hydroxide	1.43×10^{-14}
Calcium hydroxide	5.02×10^{-6}	Lead (II) phosphate	8.0×10^{-43}
Calcium phosphate	2.07×10^{-29}	Lead (II) sulfate	2.53×10^{-8}
Calcium sulfate	4.93×10^{-5}	Magnesium carbonate	6.82×10^{-8}
Copper (II) carbonate	1.4×10^{-10}	Magnesium hydroxide	5.61×10^{-12}
Copper (II) hydroxide	2.2×10^{-20}	Magnesium phosphate	1.04×10^{-24}
Copper (II) phosphate	1.4×10^{-37}	Silver bromide	5.35×10^{-13}
Iron (III) carbonate	3.13×10^{-11}	Silver chloride	1.77×10^{-10}
Iron (III) hydroxide	4.87×10^{-17}	Silver carbonate	8.46×10^{-12}
Iron (III) phosphate	2.79×10^{-39}	Silver hydroxide	2.0×10^{-8}
Iron (III) sulfate	9.91×10^{-16}	Silver iodide	8.52×10^{-17}
		Silver phosphate	8.89×10^{-17}
		Silver sulfate	1.20×10^{-5}

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Bryan Earl, L. D. R. Wilford



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