

Essential Tool for Chemistry Concepts, Variables and Equations including Sample Problems, Common Pitfalls and Helpful Hints

Pg. 1 Basic Skills, Math Review • Pg. 2 Statistics, Atomic Data, Chemical Formula Calculations • Pg. 3 Stoichiometry
Pg. 4 Gas Law Calculations, Solids and Liquids • Pg. 5 Thermodynamics, Acid-Base Calculations • Pg. 6 Equilibrium Calculations, Kinetics

Your success in chemistry depends on your ability to solve numerical problems.

BASIC SKILLS**CALCULATOR SURVIVAL**

- Become familiar with your calculator before the exam; make sure you can multiply, divide, add, subtract and use all needed functions.
- Calculators never make mistakes; they take your input (including its mistakes) and give an answer.
- Look at the answer; does it make sense?
- Do a quick estimate to check your work.

Sample: $4.341 \times 7.668 / 11.005 \times 9.035$ is roughly $4 \times 8 / 11 \times 10$ or 3.2; the actual answer: 3.24.

HOW TO DO WORD PROBLEMS

- Read and evaluate the question before you start plugging numbers into the calculator.
- Identify the variables, constants and equations.
- Write out units of the variables and constants.
- Work out the unit before the number.
- You may have extra information; or you may need to obtain constants from your text.

HOW TO WORK WITH UNITS

All numerical data has units. In chemistry we use metric "SI" units.

Partial: If the unit is wrong, the answer is wrong!

1. **Unit prefixes:** decimal powers of "10"
 - tera T 10^{12}
 - kilo K 10^3
 - milli m 10^{-3}
 - pico p 10^{-12}
2. **Fundamental Units:**
 - Mass: kilogram (kg)
 - Length: meter (m)
 - Temperature: Kelvin (K)
 - Time: second (s)
 - Amount of a substance: mole
 - Electric charge: coulomb (C)
3. **Derived Units:**
 - Area: length squared, m^2
 - Volume: length cubed, m^3 ; 1 liter (L) = 1 dm^3
 - Density: mass/volume; common unit kg/m³
 - Speed: distance/time; common unit m/s
 - Electric current: ampere (A) = 1 C/s
 - Force: Newton (N) = 1 kg m/s²
 - Energy: Joule (J) = 1 kg m/s²
 - Pressure: Pascal (Pa) = 1 N/meter²

**BASIC SKILLS cont.****HOW TO CONVERT DATA**

The unit A is converted to unit B by using a conversion factor or equation.

HOW TO USE EQUATIONS FOR DATA CONVERSION

Sample: Convert a temperature of 45°C to F and Kelvin
Given: K = $^{\circ}\text{C} + 273.15$
Calculator: Temperature in K = $45^{\circ}\text{C} + 273.15$
= 318.15 K
Given: $^{\circ}\text{F} = 9/5 \cdot ^{\circ}\text{C} + 32$
Calculator: Temperature in $^{\circ}\text{F} = 9/5 \times 45^{\circ}\text{C} + 32$
= 81 + 32 = 113 $^{\circ}\text{F}$

$$T(\text{K}) = T(\text{C}) + 273.15 \quad T(\text{F}) = 9/5 \times T(\text{C}) + 32$$

**HOW TO USE A CONVERSION FACTOR**

Sample: "1 hour = 60 minutes" gives two conversion factors:

1. Multiply by "1 hour/60 minutes" to convert minutes to hours.
2. Multiply by "60 minutes/1 hour" to convert hours to minutes.

Sample: Determine the number of hours in 45 minutes.
Given: Conversion factor is "1 hour/60 minutes."
Calculator: Time = 45 minutes $\times (1 \text{ hour}/60 \text{ minutes})$
= 0.75 hour (minutes canceled)

Common conversions:

- 1 calorie = 4.184 J
- 1 kg = 2.2 lb
- 1 m = 3.3 yd
- 1 kg = 0.9842 L
- 1 megapascal (Pa) = 1×10^{12} Pa
- 1 atm = 101325 $\times 10^2$ Pa
- 1 atm = 760 mm Hg

Partial: Equations and conversion factors have units. Be sure of converting conversion factors.

MATH REVIEW**HOW TO WORK WITH ALGEBRAIC EQUATIONS**

Give equal treatment to each side.

- Add or subtract:

Sample: Given: $x = y$, then $x + z = y + z$.
- Multiply or divide:

Sample: Given: $x = y$, then $4x = 4y$ and $x/3 = y/3$.

Partial: Division by zero is not allowed.

HOW TO WORK WITH SCIENTIFIC NOTATION

The exponent gives the power of 10.

Sample: $0.000045 = 4.5 \times 10^{-5}$; $1345 = 1.345 \times 10^3$

Sample: chemical applications: Molecular dimensions are 10^{-10} m; 1 liter of water contains about 10^{23} atoms.

MATH REVIEW cont.**HOW TO USE LOGARITHMS & EXPONENTS**

Common Logarithms: \log_{10} base 10

Decays number or function in powers of 10.

Sample: Given: $y = 10^{-x}$, then $\log_{10} y = -x$.

Natural Logarithms: \ln base e $e = 2.718281828$.

Decays number or function in powers of "e".

Sample: Given: $x = e^y$, then $\ln x = y$.

- Product: $\log xy = \log x + \log y$
- Powers: $\log x^y = y \log x$
- Multiplication: Add exponents from each term.
- Sample:** $10^3 \times 10^7 = 10^{3+7} = 10^{10}$
- Division: Subtract denominator exponents from numerator exponents.
- Sample:** $10^7/10^3 = 10^{7-3} = 10^4$

• Square root: $\sqrt{a} = a^{1/2}$

• Cubes: $\sqrt[3]{a} = a^{1/3}$

Partial: Your calculator has separate keys for \log (base 10), \ln (base e), and e^x .

Sample: chemical applications: pH of acid and base.

HOW TO CALCULATE ROOTS OF A POLYNOMIAL

An equation of the form $ax^2 + bx + c = 0$, has 2 solutions or roots, given by the quadratic formula:

$$\frac{-b + \sqrt{b^2 - 4ac}}{2a} \quad \text{Root 1}$$

$$\frac{-b - \sqrt{b^2 - 4ac}}{2a} \quad \text{Root 2}$$

Sample: Determine the roots for the equation $3x^2 + 4x + 1 = 0$.

$$\text{Given: } a = 3, b = 4, c = 1$$

$$\text{Calculator: Root 1} = [-4 + \sqrt{16 - 4(3)(1)}]/(2(3))$$

$$= [-4 + \sqrt{16 - 12}]/6 = [-4 + 2]/6 = -1/3$$

$$\text{Calculator: Root 2} = [-4 - \sqrt{16 - 12}]/6 = [-4 - 2]/6 = -1$$

Partial: Beware round-off error. Substitute the roots into the equation to verify results.

Sample: chemical applications: Weak acids, weak bases, buffers, chemical equilibrium

HOW TO DETERMINE THE EQUATION OF A LINE**LINEAR EQUATION: $y = mx + b$**

• Slope of the line: $m = \Delta y/\Delta x$

In a line graph, the line crosses the y-axis at "b"; $b = y_1 - mx_1$

Sample: Determine the equation of a line using these data:

$$\text{Given: } x_1 = 1, y_1 = 2, x_2 = 3, y_2 = 3$$

$$\text{Calculator: Slope} = m = \Delta y/\Delta x$$

$$= (3 - 2)/(3 - 1) = 1/2 = 0.5$$

$$\text{Calculator: } y = \text{Intercept} = b = y_1 - mx_1$$

$$= (2 - 0.5)(1) = 1.5 = 1.5$$

The equation of the line is $y = 0.5x + 1.5$.

Sample: chemical applications: Gas calculations; stoichiometry; analyzing reaction-rate data



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David Kirk

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Chemistry Equations and Answers Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students **Pp/Chemistry** BarCharts, Inc.,2008-06-18

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