

Chemistry 100 Final Study Guide Key

- **Chapter 1: Introduction**
 - a. Physical vs. Chemical properties
 - 1) Physical properties: shiny, metallic, malleable, ductile, conductive. Chemical property: reactive with water.
 - b. Units of measurement
 - 1) Meters, liters, grams (with any prefixes) are metric system. All others are U.S.
 - 2) 5.68×10^{-7} = 46,300,000
 - c. Density
 - 1) 2.47 g/mL
- **Chapter 3: Atomic Structure**
 - a. Historical Perspective
 - 1) List the major scientific accomplishment of the following scientists:
 - (i) Experiments in electrochemistry described ions and the electrical nature of atoms.
 - (ii) Cathode-ray tube experiments led to the proposal of the "Plum pudding" model of atoms.
 - (iii) Oil-drop experiments calculated the mass of electrons.
 - (iv) Gold foil experiments proved atoms had a dense, positively-charged nucleus and empty space.
 - (v) Proposed an atomic model with a nucleus and electrons circling like planets.
 - (vi) Proposed that electrons exist in a "cloud" or probability region around the nucleus.
 - b. Isotopes
 - 1) Isotopes of an element have the same number of protons, but different numbers of neutrons. Regular hydrogen has 1 proton and no neutrons. Isotopes could have 1 proton and 1 neutron (deuterium) or 1 proton and 2 neutrons (tritium).
 - c. Structure of the atom
 - 1) Protons and neutrons are in the nucleus. Electrons are in regions of space surrounding the nucleus.
 - 2) 3 energy levels, 4 sublevels, 6 orbitals, 11 electrons
 - 3) $1s^2 2s^2 2p^6 3s^2 3p^2$
 - d. Electron configurations and trends on the Periodic Table
 - 1) See the periodic Table.
- **Chapter 5: Chemical bonds**
 - a. Electron-dot structures
 - 1) Draw the electron-dot structure of an oxygen atom.
 - b. Octet Rule and Ions
 - 1) Mg^{2+} , Cl^- , N^{3-}
 - c. Ionic compounds
 - 1) Aluminum bromide, calcium sulfide, sodium nitride
 - d. Covalent compounds
 - 1) Sulfur dioxide, phosphorus trichloride, carbon tetrahydride
 - e. Polyatomic molecules
 - 1) Nitrate (NO_3^-), chlorate (ClO_3^-), carbonate (CO_3^{2-}), sulfate (SO_4^{2-}), chromate (CrO_4^{2-}), phosphate (PO_4^{3-})
 - f. Electron-dot formulas for molecules
 - 1) Answer will be given in class.
 - g. VSEPR Model and Molecular shape



VSEPR shape = pyramidal



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Sharon Sherman, Alan Sherman

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