

# Chemistry 100 Final Study Guide Key

- **Chapter 1: Introduction**

- a. Physical vs. Chemical properties

- 1) Physical properties: shiny, metallic, malleable, ductile, conductive. Chemical property: reactive with water.

- b. Units of measurement

- 1) Meters, liters, grams (with any prefixes) are metric system. All others are U.S.

- 2)  $5.68 \times 10^{-7}$       46,300,000

- c. Density

- 1) 2.47g/mL

- **Chapter 3: Atomic Structure**

- a. Historical Perspective

- 1) List the major scientific accomplishment of the following scientists:

- (i) Experiments in electrochemistry described ions and the electrical nature of atoms

- (ii) Cathode-ray tube experiments led to the proposal of the "Plum pudding" model of atoms

- (iii) Oil-drop experiments calculated the mass of electrons

- (iv) Gold foil experiments proved atoms had a dense, positively-charged nucleus and empty space

- (v) Proposed an atomic model with a nucleus and electrons circling like planets

- (vi) Proposed that electrons exist in a "cloud" or probability region around the nucleus

- b. Isotopes

- 1) Isotopes of an element have the same number of protons, but different numbers of neutrons. Regular hydrogen has 1 proton and no neutrons. Isotopes could have 1 proton and 1 neutron (deuterium) or 1 proton and 2 neutrons (tritium).

- c. Structure of the atom

- 1) Protons and neutrons are in the nucleus. Electrons are in regions of space surrounding the nucleus.

- 2) 3 energy levels, 4 sublevels, 6 orbitals, 11 electrons

- 3)  $1s^2 2s^2 2p^6 3s^2 3p^2$

- d. Electron configurations and trends on the Periodic Table

- 1) See the periodic Table.

- **Chapter 5: Chemical bonds**

- a. Electron-dot structures

- 1) Draw the electron-dot structure of an oxygen atom.



- b. Octet Rule and Ions

- 1)  $\text{Mg}^{2+}$ ,  $\text{Cl}^-$ ,  $\text{N}^{3-}$

- c. Ionic compounds

- 1) Aluminum bromide, calcium sulfide, sodium nitride

- d. Covalent compounds

- 1) Sulfur dioxide, phosphorus trichloride, carbon tetrahydride

- e. Polyatomic molecules

- 1) Nitrate ( $\text{NO}_3^-$ ), chlorate ( $\text{ClO}_3^-$ ), carbonate ( $\text{CO}_3^{2-}$ ), sulfate ( $\text{SO}_4^{2-}$ ), chromate ( $\text{CrO}_4^{2-}$ ), phosphate ( $\text{PO}_4^{3-}$ )

- f. Electron-dot formulas for molecules

- 1) Answer will be given in class.

- g. VSEPR Model and Molecular shape



VSEPR shape = pyramidal

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**Sharon Sherman, Alan Sherman**



## **Chemistry Final Study Guide Georgia Answers:**

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