

Chemistry 100 Final Study Guide

- **Chapter 1: Introduction**

- a. Physical vs. Chemical properties
 - 1) List 2 physical and 1 chemical property of magnesium.
- b. Units of measurement
 - 1) Can you distinguish between metric and U.S. units of measurement?
 - 2) What is 0.000000568 in scientific notation? What is 4.63×10^7 in regular notation?
- c. Density
 - 1) What is the density of a 37g block of metal that takes up 15mL of volume?

- **Chapter 3: Atomic Structure**

- a. Historical Perspective
 - 1) List the major scientific accomplishments of the following scientists:
 - (i) Michael Faraday
 - (ii) Joseph Thomson
 - (iii) Robert Millikan
 - (iv) Ernest Rutherford
 - (v) Neils Bohr
 - (vi) Erwin Schrodinger
- b. Isotopes
 - 1) Define an isotope and give an example of several isotopes for hydrogen.
- c. Structure of the atom
 - 1) Name the 3 main parts of an atom, and indicate where in the atom they'd be found.
 - 2) For the element sodium: how many energy levels? How many sublevels? What orbitals? How many electrons?
 - 3) Write the full electron configuration of silicon.
- d. Electron configurations and trends on the Periodic Table
 - 1) Locate each of the following on the Periodic Table: alkali metals, alkaline earth metals, transition metals, halogens, noble gases, metals, non-metals.

- **Chapter 5: Chemical bonds**

- a. Electron-dot structures
 - 1) Draw the electron-dot structure of an oxygen atom.
- b. Octet Rule and Ions
 - 1) What charge of ions do magnesium, chlorine and nitrogen form?
- c. Ionic compounds
 - 1) Name the following compounds: AlBr_3 , CaS , Na_3N
- d. Covalent compounds
 - 1) Name the following compounds: SO_2 , PCl_3 , CH_4
- e. Polyatomic molecules
 - 1) Name and write the formula for the 6 -ate polyatomic molecules we discussed.
- f. Electron-dot formulas for molecules
 - 1) Write the electron-dot formula for each of the 6 -ate polyatomic ions we discussed.
- g. VSEPR Model and Molecular shape
 - 1) Draw the electron-dot formula of ammonia (NH_3) and state its VSEPR shape.
- h. Molecule polarity
 - 1) Which of the VSEPR molecule shapes are polar and which are non-polar?

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Birhanu Aberha Berele



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League Nursing, 2008-09-12 Providing an overview of the math science and reading comprehension skills necessary for admission to LPN programs this bestselling guide includes multiple choice questions and practice examinations in each of the three test areas **5 Steps to a 5 AP Chemistry, 2010-2011 Edition** John Moore, Richard H. Langley, 2010-01-01 A

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