

## Chemistry First Nine Weeks study guide for exam

1. What are SI base units for time, length, amount of substance, mass?
2. Express numbers in scientific notation given standard, and given scientific put in standard.  
 $452.7 \text{ m} = ?$ ,  $0.00077893 \text{ L} = ?$ ,  $6.77 \times 10^4 \text{ mL} = ?$ ,  $8.12 \times 10^{-6} \text{ km} = ?$
3. What are definitions of precision and accuracy and be able to apply them correctly to data given. See data tables in chapter 2 for examples.
4. Significant digits- how many in a given number, which types of zeroes are significant.  
ex.  $0.005603 \text{ m}$ ,  $1241 \text{ m}$ ,  $4.5800500 \text{ cm}$ ,  $23.0 \text{ L}$
5. Correct rounding with significant digits.  $4.5677 \text{ L}$  to 3 sig. dig,  $22.8655$  to 4 sig.dig.
6. Adding, subtracting, multiplying, dividing correctly with sig. digits. (2 SEPARATE RULES)  
 $34.41 \text{ m} + 11.8764 \text{ m} = ?$ ,  $56.34 \text{ m} - 21.9 \text{ m} = ?$ ,  $439 \text{ cm} \times 5.33 \text{ cm} = ?$ ,  $189.22 \text{ g} / 76 \text{ m}$
7. Conversion of SI units to each other using dimensional analysis. You are responsible for the units you were given in class. Also make sure you know nano. Ex:  $10^9 \text{ nm} = 1 \text{ m}$ , very small unit, new branch of chemistry springing up around thee very tiny! You will hear more of this in future.
8. Density problems. Find the third term if given the other 2. Know the formula; it won't be given. Find density of a rock with mass of  $35.8 \text{ g}$  and volume of  $8.1 \text{ mL}$ . Find the volume of a substance with mass of  $163 \text{ grams}$  and density of  $5.77 \text{ g/mL}$ .
9. Dimensional analysis problems. See review problem sets for examples.
10. Characteristics of the 3 basic states of matter as they pertain to their container, volume, shape, etc.
11. Recognition of physical properties and chemical properties and/or changes. Copper carbonate which is green in color, when heated, turns black and loses mass. Grinding charcoal into a fine powder is an example of which?
12. What principles and means are used to separate mixtures?
13. Distinguish between homogeneous and heterogeneous mixtures. Well mixed koolaid, tossed salad, unsweetened iced tea, orange juice with pulp, solution.
14. Law of Conservation of Mass. State it and explain how you can use it to solve basic problems. Also % mass of an element in a compound. See practice problem sets.
15. State 4-5 parts of Dalton's atomic theory. Whose idea had he revived? Which part(s) do we not agree with completely anymore?
16. Three subatomic particles in an atom and where they are located. Be able to find all three of them if given atomic mass (and PT). Know how to write both hyphen notation and nuclear notation.
16. What were the characteristics of the nucleus that Rutherford discovered? What experiment did he discover them in?

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**R Sanford**



## Chemistry First Nine Weeks Study Guide:

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## **Table of Contents Chemistry First Nine Weeks Study Guide**

1. Understanding the eBook Chemistry First Nine Weeks Study Guide
  - The Rise of Digital Reading Chemistry First Nine Weeks Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry First Nine Weeks Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry First Nine Weeks Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry First Nine Weeks Study Guide
  - Personalized Recommendations
  - Chemistry First Nine Weeks Study Guide User Reviews and Ratings
  - Chemistry First Nine Weeks Study Guide and Bestseller Lists
5. Accessing Chemistry First Nine Weeks Study Guide Free and Paid eBooks
  - Chemistry First Nine Weeks Study Guide Public Domain eBooks
  - Chemistry First Nine Weeks Study Guide eBook Subscription Services
  - Chemistry First Nine Weeks Study Guide Budget-Friendly Options

6. Navigating Chemistry First Nine Weeks Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry First Nine Weeks Study Guide Compatibility with Devices
  - Chemistry First Nine Weeks Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry First Nine Weeks Study Guide
  - Highlighting and Note-Taking Chemistry First Nine Weeks Study Guide
  - Interactive Elements Chemistry First Nine Weeks Study Guide
8. Staying Engaged with Chemistry First Nine Weeks Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry First Nine Weeks Study Guide
9. Balancing eBooks and Physical Books Chemistry First Nine Weeks Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry First Nine Weeks Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry First Nine Weeks Study Guide
  - Setting Reading Goals Chemistry First Nine Weeks Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry First Nine Weeks Study Guide
  - Fact-Checking eBook Content of Chemistry First Nine Weeks Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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