

As Mendeleev experimented with a variety of different elements, he began to notice a trend associated with certain types of elements. Let's use lithium, potassium, and sodium as examples. These elements are silver-colored, don't exist by themselves in nature, and tend to be very reactive. When these elements react with oxygen, they react in the same whole number ratio of atoms. For example, lithium reacts with oxygen to form Li_2O . Potassium and sodium will react in the same whole number ratio, generating K_2O and Na_2O respectively. This was not just a coincidence. Mendeleev experimented with other elements and found similar patterns for other metals as well. Magnesium, calcium and beryllium all react with oxygen in a 1:1 ratio (i.e. BeO , MgO , and CaO). There is a noticeable repeating pattern of chemical and physical properties every 8 elements. Notice similarities to the periodic table of Mendeleev and the modern day periodic table. They all exhibit very similar patterns in regards to their reactivity. Found in groups.

Periodic Table of Elements Based on Mendeleev's Periodic Law																	
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H	2 He	3 Li	4 Be	5 B	6 C	7 N	8 O	9 F	10 Ne	11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	19 F	20 Ne	21 Na	22 Mg	23 Al	24 Si	25 P	26 S	27 Cl	28 Ar
29 K	30 Ca	31 Sc	32 Ti	33 V	34 Cr	35 Mn	36 Fe	37 Co	38 Ni	39 Cu	40 Zn	41 Ga	42 Ge	43 As	44 Se	45 Br	46 Kr
47 Rb	48 Sr	49 Y	50 Zr	51 Nb	52 Ta	53 Hf	54 Ta	55 W	56 Re	57 Os	58 Ru	59 Rh	60 Os	61 Ir	62 Pt	63 Au	64 Hg
65 Cs	66 Ba	67 La	68 Ce	69 Pr	70 Nd	71 Pm	72 Sm	73 Eu	74 Gd	75 Tb	76 Dy	77 Ho	78 Er	79 Tm	80 Yb	81 Lu	82 Lu
83 Lu	84 Ce	85 Pr	86 Nd	87 Pm	88 Sm	89 Eu	90 Gd	91 Tb	92 Dy	93 Ho	94 Er	95 Tm	96 Yb	97 Lu	98 Lu	99 Lu	100 Lu

Even more amazingly, he left gaps in the periodic table where elements had not been discovered yet. Ge (germanium) had not been discovered, yet based on the trends on the periodic table, he was able to predict many of the physical properties of Ge, such as atomic mass, density, melting point, and its formula when reacting with oxygen.

5.2 - The Modern Periodic Table

The modern day periodic table is quite elegant and simple in design. Each vertical column on the periodic table is known as a **group** and each horizontal row is known as a **period**. Both groups and rows are numbered. Sometimes the groups can be labeled a bit differently depending on the periodic table which you are using. Groups are labeled with either an A or a B after the roman numeral. Groups 1A-8A are known as the **main group elements**. Groups 1B-8B are known as the **transition elements**. The two rows at the bottom are known as **lanthanides** and **actinides** respectively. Main group elements are what most of early chemistry was based upon. These are the elements the rule of 8 applies to in our previous part of this unit. Many of the transition metals have very similar properties. The numbering system can look a bit odd, especially when looking at the transition metals. 8 groups have the title **VIIIb**. Don't go making an absolute casserole out of this concept.

Periodic Table of the Elements																	
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H	2 He	3 Li	4 Be	5 B	6 C	7 N	8 O	9 F	10 Ne	11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	19 F	20 Ne	21 Na	22 Mg	23 Al	24 Si	25 P	26 S	27 Cl	28 Ar
29 K	30 Ca	31 Sc	32 Ti	33 V	34 Cr	35 Mn	36 Fe	37 Co	38 Ni	39 Cu	40 Zn	41 Ga	42 Ge	43 As	44 Se	45 Br	46 Kr
47 Rb	48 Sr	49 Y	50 Zr	51 Nb	52 Ta	53 Hf	54 Ta	55 W	56 Re	57 Os	58 Ru	59 Rh	60 Os	61 Ir	62 Pt	63 Au	64 Hg
65 Cs	66 Ba	67 La	68 Ce	69 Pr	70 Nd	71 Pm	72 Sm	73 Eu	74 Gd	75 Tb	76 Dy	77 Ho	78 Er	79 Tm	80 Yb	81 Lu	82 Lu
83 Lu	84 Ce	85 Pr	86 Nd	87 Pm	88 Sm	89 Eu	90 Gd	91 Tb	92 Dy	93 Ho	94 Er	95 Tm	96 Yb	97 Lu	98 Lu	99 Lu	100 Lu

In addition to groups and periods, we find that the elemental arrangement of the periodic table also

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