

## Ionic Compounds and Metals

### Section 7.1 Ion Formation

pages 206–209

#### Section 7.1 Assessment

page 209

- 1. Compare** the stability of a lithium atom with that of its ion,  $\text{Li}^+$ .

The  $\text{Li}^+$  ion is more stable because it has a complete octet.

- 2. Describe** two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.

- 3. Apply** Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.

- 4. Summarize** ionic bond formation by correctly pairing these terms: *cation*, *anion*, *electron gain*, *electron loss*.

(anion, electron gain), (cation, electron loss)

- 5. Apply** Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.

- nitrogen

$[\text{He}]2s^22p^3$ ; gain 3 electrons ( $3^-$  ion) or lose 5 electrons ( $5^+$  ion)

- sulfur

$[\text{Ne}]3s^23p^4$ ; gain 2 electrons ( $2^-$  ion)

- barium

$[\text{Xe}]6s^2$ ; lose 2 electrons ( $2^+$  ion)

- lithium

$[\text{He}]2s^1$ ; lose 1 electron ( $1^+$  ion)

- 6. Model** Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming  $\text{Ca}^{2+}$ , and that bromine gains one electron, forming  $\text{Br}^-$ . The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

### Section 7.2 Ionic Bonds and Ionic Compounds

pages 210–217

#### Practice Problems

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Explain how an ionic compound forms from these elements.

- sodium and nitrogen

Three Na atoms each lose  $1 e^-$ , forming  $1^+$  ions. One N atom gains  $3 e^-$ , forming a  $3^-$  ion. The ions attract, forming  $\text{Na}_3\text{N}$ .

$$3 \text{ Na ions } \left( \frac{1+}{\text{Na ion}} \right) + 1 \text{ N ion } \left( \frac{3-}{\text{N ion}} \right) \\ = 3(1+) + 1(3-) = 0$$

The overall charge on one formula unit of  $\text{Na}_3\text{N}$  is zero.

- lithium and oxygen

Two Li atoms each lose  $1 e^-$ , forming  $1^+$  ions. One O atom gains  $2 e^-$ , forming a  $2^-$  ion. The ions attract, forming  $\text{Li}_2\text{O}$ .

$$2 \text{ Li ions } \left( \frac{1+}{\text{Li ion}} \right) + 1 \text{ O ion } \left( \frac{2-}{\text{O ion}} \right) \\ = 2(1+) + 1(2-) = 0$$

The overall charge on one formula unit of  $\text{Li}_2\text{O}$  is zero.

# Chemistry Matter And Change Chapter 7 Study Guide Answers

**John McMurry, Susan McMurry**



## **Chemistry Matter And Change Chapter 7 Study Guide Answers:**

*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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