

AP Chemistry Pacing Guide

- Approximation - Multistep Reaction - Energy Profile - Catalysis		first step is not rate limiting. 10. I can represent the activation energy and overall energy change in a multistep reaction with a reaction energy profile. 11. I can explain the relationship between the effect of a catalyst on a reaction and changes in the reaction mechanism.	
Unit 6: Thermodynamics - Endothermic and Exothermic Processes - Energy Diagrams - Heat Transfer and Thermal Equilibrium - Heat Capacity and Calorimetry - Energy of Phase Changes - Introduction to Enthalpy of Reaction - Bond Enthalpies - Enthalpy of Formation - Hess's Law	BIG IDEA 4: Energy	- I can explain the relationship between experimental observations and energy changes associated with a chemical or physical transformation. - I can represent a chemical or physical transformation with an energy diagram. - I can explain the relationship between the transfer of thermal energy and molecular collisions. - I can calculate the heat q absorbed or released by a system undergoing heating/cooling based on the amount of the substance, the heat capacity, and the change in temperature. - I can explain changes in the heat absorbed or released by a system undergoing a phase transition based on the amount of the substance in moles and the molar enthalpy of the phase transition. - I can calculate the heat q absorbed or released by a system undergoing a chemical reaction in relation to the amount of the reacting substance in moles and the molar enthalpy of reaction. - I can calculate the enthalpy change of a reaction based on the average bond energies of bonds broken and formed in the reaction. - I can calculate the enthalpy change for a chemical or physical process based on the standard enthalpies of formation. - I can explain the relationship between the enthalpy of a chemical or physical process and the sum of the enthalpies of the individual steps.	10 days
Unit 7: Equilibrium - Introduction to Equilibrium - Direction of Reversible Reactions - Reaction Quotient and Equilibrium Constant - Calculating the Equilibrium Constant - Magnitude of the Equilibrium Constant	BIG IDEA 1: Scale, Proportion, and Quantity BIG IDEA 3: Transformation	- I can explain the relationship between the direction in which a reversible reaction proceeds and the relative rates of the forward and reverse reactions. - I can represent the reaction quotient Q_c or Q_p for a reversible reaction, and the corresponding equilibrium expressions $K_c = Q_c$ or $K_p = Q_p$. - I can calculate K_c or K_p based on experimental observations of concentrations or pressures at equilibrium. - I can explain the relationship between very large or very small values of K and the relative concentrations of chemical species at equilibrium. - I can represent a multistep process with an overall equilibrium expression, using the constituent K expressions for each individual	9 days

Chemistry Pacing Guide

Henry Robinson Luce



Chemistry Pacing Guide:

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