

REDOX REACTION

Solved Examples

Example - 6 :



Base Acid

Sol. Valency factor of base = 1

Here, two molecule of NaOH replaced 2H^+ ion from the H_2SO_4 . Therefore, each molecule of NaOH replaced only one H^+ ion of acid, so v.f. = 1.

○ v. f. for acid is the number of OH^- replaced from the base by each molecule of acid

Solved Examples

Example - 7 :



Base Acid

Sol. Valency factor of acid = 1

Here, one of molecule of H_2SO_4 replaced one OH^- from NaOH. Therefore, valency factor for H_2SO_4 is one

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

○ Salts :

(a) In non-reacting condition

○ Valency factor = Total number of positive charge or negative charge present in the compound.

Solved Examples

Example - 8 :

		Na_2CO_3	$\text{Fe}_2(\text{SO}_4)_3$	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
Sol.	v.f.	2	$2 \times 3 = 6$	2
	Eq. wt.	$\frac{M}{2}$	$\frac{M}{6}$	$\frac{M}{2}$

(b) In reacting condition

Solved Examples

Example - 9 :



Base Acid

Sol. It is an acid base reaction, therefore valency factor for Na_2CO_3 is one while in non-reacting condition, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

Chemistry Redox Reactions Study Guide Answers

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Table of Contents Chemistry Redox Reactions Study Guide Answers

1. Understanding the eBook Chemistry Redox Reactions Study Guide Answers
 - The Rise of Digital Reading Chemistry Redox Reactions Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Redox Reactions Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Redox Reactions Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Redox Reactions Study Guide Answers
 - Personalized Recommendations
 - Chemistry Redox Reactions Study Guide Answers User Reviews and Ratings
 - Chemistry Redox Reactions Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Redox Reactions Study Guide Answers Free and Paid eBooks
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 - Chemistry Redox Reactions Study Guide Answers eBook Subscription Services
 - Chemistry Redox Reactions Study Guide Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
- Chemistry Redox Reactions Study Guide Answers Compatibility with Devices
- Chemistry Redox Reactions Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Redox Reactions Study Guide Answers
 - Highlighting and Note-Taking Chemistry Redox Reactions Study Guide Answers
 - Interactive Elements Chemistry Redox Reactions Study Guide Answers
- 8. Staying Engaged with Chemistry Redox Reactions Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Redox Reactions Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Redox Reactions Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Redox Reactions Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Redox Reactions Study Guide Answers
 - Setting Reading Goals Chemistry Redox Reactions Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Redox Reactions Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Redox Reactions Study Guide Answers
 - Distinguishing Credible Sources
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 - Interactive and Gamified eBooks

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