

REDOX REACTION

Solved Examples

Example - 6 :



Sol. Valency factor of base = 1

Here, two molecule of NaOH replaced 2H^+ ion from the H_2SO_4 . Therefore, each molecule of NaOH replaced only one H^+ ion of acid, so v.f. = 1.

Q. v. f. for acid is the number of OH^- replaced from the base by each molecule of acid.

Solved Examples

Example - 7 :



Sol. Valency factor of acid = 1

Here, one molecule of H_2SO_4 replaced one OH^- from NaOH. Therefore, valency factor for H_2SO_4 is one.

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

Q. Salts :

(a) In non-reacting condition

Q. Valency factor = Total number of positive charge or negative charge present in the compound.

Solved Examples

Example - 8 :

Sol.	Na_2CO_3	$\text{Fe}_2(\text{SO}_4)_3$	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
	v.f.	2	2
Eq. wt.	$\frac{M}{2}$	$\frac{M}{6}$	$\frac{M}{2}$

(b) In reacting condition

Solved Examples

Example - 9 :



Sol. It is an acid base reaction, therefore valency factor for Na_2CO_3 is one while in non-reacting conditions, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

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