



Topic 12 – Atomic structure (AHL)

End of topic questions [page 300]

1. B: when solving this type of question, examine the graph or tabulated data to determine the point at which there is a rapid increase in the ionisation energy; the largest jump occurs between IE_5 and IE_6 ; this corresponds to a change in energy level and therefore the element must be in Group 14.
2. C: Boron is found in group 13; it will lose 5 successive valence electrons before a change in energy level takes place; therefore the greatest difference is between IE_5 and IE_6 .
3. C: answer A is incorrect because successive ionisation energy values increase, not decrease; answer B is incorrect because nuclear ionisation energy are measured in kJ mol^{-1} ; alternative D is incorrect because ionisation energy increase (not decrease) across the period as effective nuclear charge increases and atomic radii decrease.
4. B: first ionisation energy: $\text{M(g)} \rightarrow \text{M}^+(\text{g}) + e^-$ the maximum energy (in kJ mol^{-1}) to remove one electron from a gaseous atom/the energy required to remove one mole of electrons from one mole of gaseous atoms; periodicity: repeating patterns of physical and chemical properties
 Wp evidence for main levels: highest values for noble gases/lowest values for alkali metals; general increase across a period; evidence for sub-levels: dip in I.E. from the 1s/2s to 3s/Mg to 4s/Group 2 to Group 3; dip in I.E. from N to O/P to 5s/Group 5 to Group 6;
 BQ: $\text{M}^+(\text{g}) \rightarrow \text{M}^{2+}(\text{g}) + e^-$

Chemistry Review Topic 12 Answers

Joacim Rocklöv



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