

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Data -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a mole ratio ratio to convert between moles of different substances.

Skills - Using a balanced equation, calculate using mole to mole ratios. Using a balanced equation, given the grams of reactant, calculate the grams of product.

- Consider the following unbalanced equation.



There are 5.00g of manganese(IV) oxide present. calculate the grams of chlorine gas.
 $5.00\text{g MnO}_2 \times 1\text{mol MnO}_2 \times 1\text{mol Cl}_2 \times 70.9\text{g Cl}_2 = 45.45\text{g Cl}_2$
 $55.94\text{g MnO}_2 \cdot 1\text{mol MnO}_2 \cdot 1\text{mol Cl}_2$

Skills - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Since it is incomplete, you calculate that a reaction should produce 4.0g Cu(OH)₂. After the lab is over you find that 3.8g Cu(OH)₂ were collected. What would the percent yield be for that reaction completed within the lab?

$$\text{% yield} = \frac{\text{Actual/Theoretical}}{100} \times 100$$

$$= \frac{3.8\text{g}}{4.0\text{g}} \times 100 = 95\%$$

Skills - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate reacts with aqueous hydrochloric acid to produce solid calcium chloride, carbon dioxide, and water vapor. What is the percent yield if 5.00g of calcium carbonate reacts with 25.0g of hydrochloric acid and only 3.42g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$1.75\text{g CaCO}_3 \times 1\text{mol CaCO}_3 / 100.09\text{g CaCO}_3 = 1.75\text{mol CaCO}_3 / 1\text{mol HCl} = 1.75\text{ mol}$$

$$25.0\text{g HCl} \times 1\text{mol HCl} / 36.45\text{g HCl} = 6.857 \text{ mol HCl} / 2\text{mol HCl} = 3.43 \text{ mol}$$

$$1.75\text{mol CaCO}_3 \times (1\text{mol CaCl}_2 / 1\text{mol CaCO}_3) \times (110.9\text{g CaCl}_2 / 1\text{mol CaCl}_2) = 1.724\text{g}$$

$$\text{% yield} = \frac{\text{Actual/Theoretical}}{100} \times 100$$

$$= \frac{1.724\text{g}}{3.43\text{g}} \times 100 = 50.05\%$$

Chemistry Semester 2 Course Review Answers Unit 9

Stoichiometry

Minjie Lin

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The Software Encyclopedia 2001 ,2001

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