

Chemistry

Stoichiometry Reflect #2

Learning Target: I can use stoichiometry to calculate amounts.

Name Ms. Hettie

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left(\frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = 1.06 \text{ moles (NH}_4\text{)}_2\text{O}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = 0.706 \text{ moles (NH}_4\text{)}_3\text{N}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left(\frac{48.16 \text{ g}}{1 \text{ mol B}} \right) = 48.16 \text{ g (NH}_4\text{)}_3\text{N}$$

2. Using the following chemical reaction: $2 \text{ Hg} + \text{Cu}(\text{NO}_3)_2 \rightarrow 2 \text{ HgNO}_3 + \text{Cu} \xrightarrow{A}$
and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left(\frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = 1.17 \times 10^{24} \text{ mc HgNO}_3$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left(\frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left(\frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

Theoretical

$$\begin{aligned} \text{yield} &= \frac{7.75 \text{ g}}{8.46 \text{ g}} \\ &= 91.4\% \end{aligned}$$

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Chemistry Stoichiometry Practice Answers

Stanley R. Radel, Marjorie H. Navidi

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Physical Chemistry John S. Winn, 1995 This text presents physical chemistry as a coherent whole rather than a set of disjointed topics and shows how the subject relates to the rest of chemistry and physics It emphasizes physical models as well as mathematical techniques along with both rigorous and approximate order of magnitude problem solving Designed to progress beyond a numerical answer problems expose the physical significance of the situation and teach students how to pose a problem in the first place In addition modern molecular concepts currently unanswered problems in research experimental techniques and new directions in the field are introduced wherever appropriate An orderly progression of thermodynamics carefully builds students knowledge without covering too much too early on Chemical reaction thermodynamics is covered in Chapter 7 after the culmination of thermodynamics with advanced material in Chapter 10

Chemistry, Problem-Solving Worktext James E. Brady, Frederick Senese, 2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and

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